

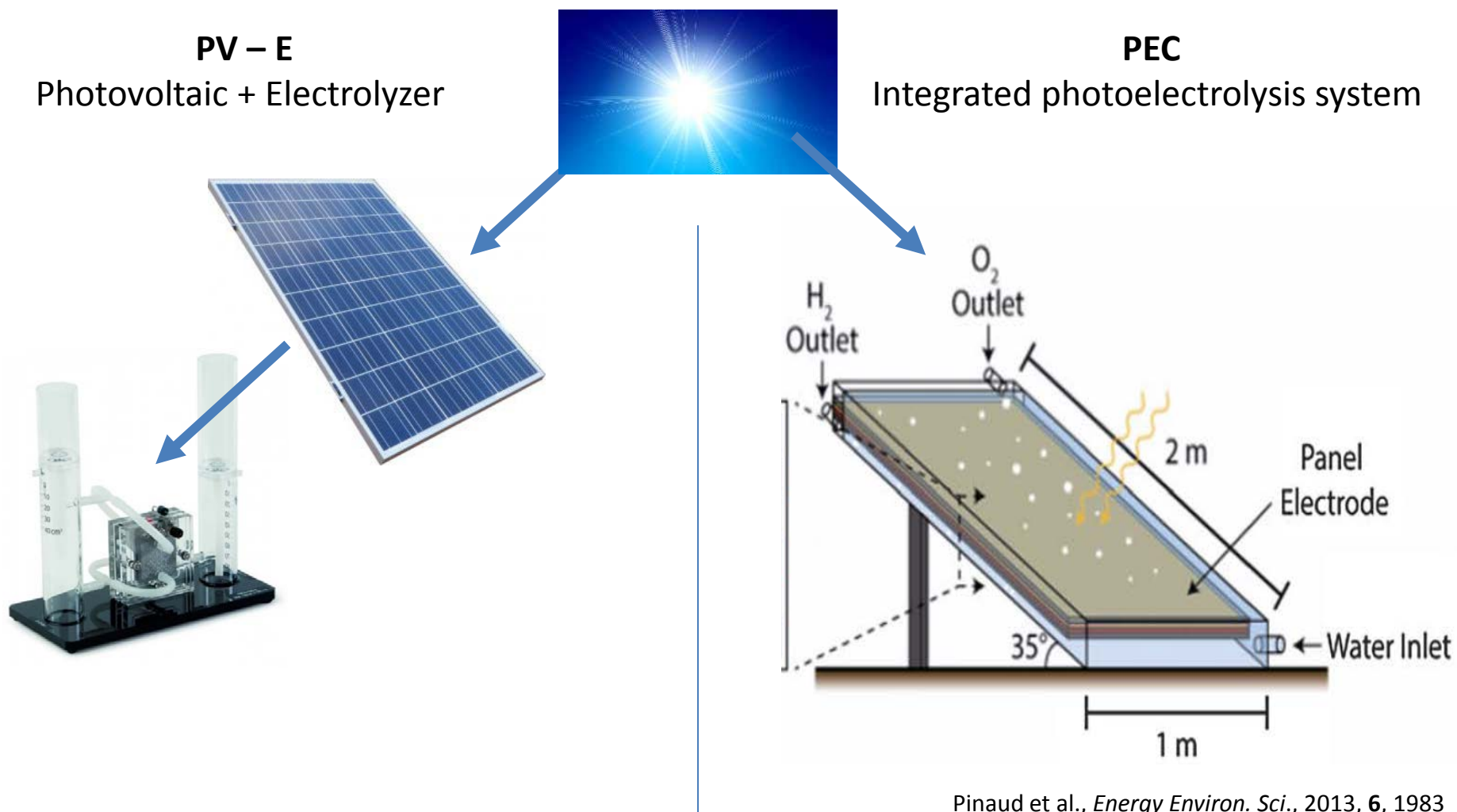
# ***MIS Photocathodes for Solar Fuel Generation***



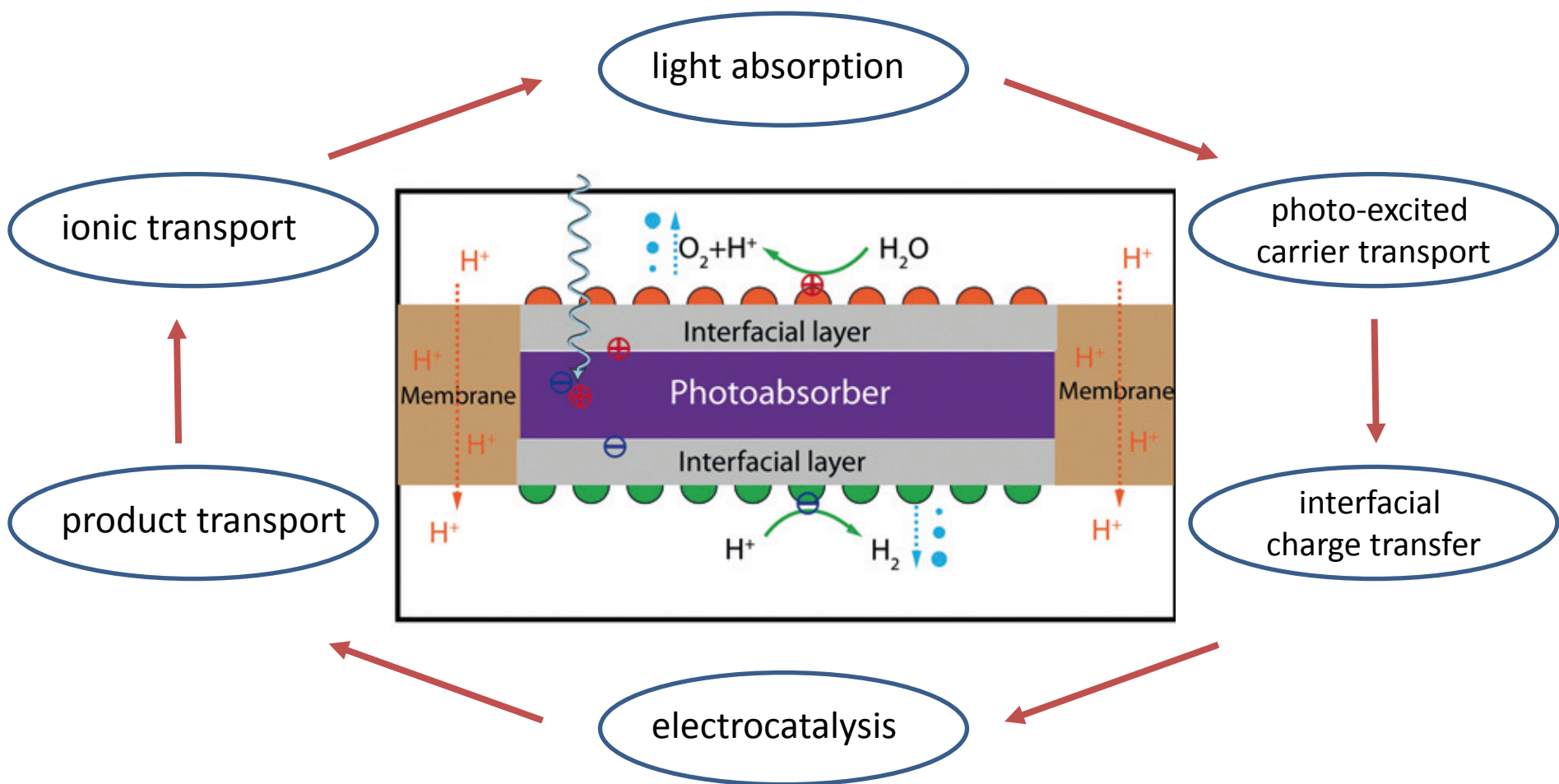
***Katharina Krischer  
Physik-Department  
TU München***



# Solar to Chemical Energy

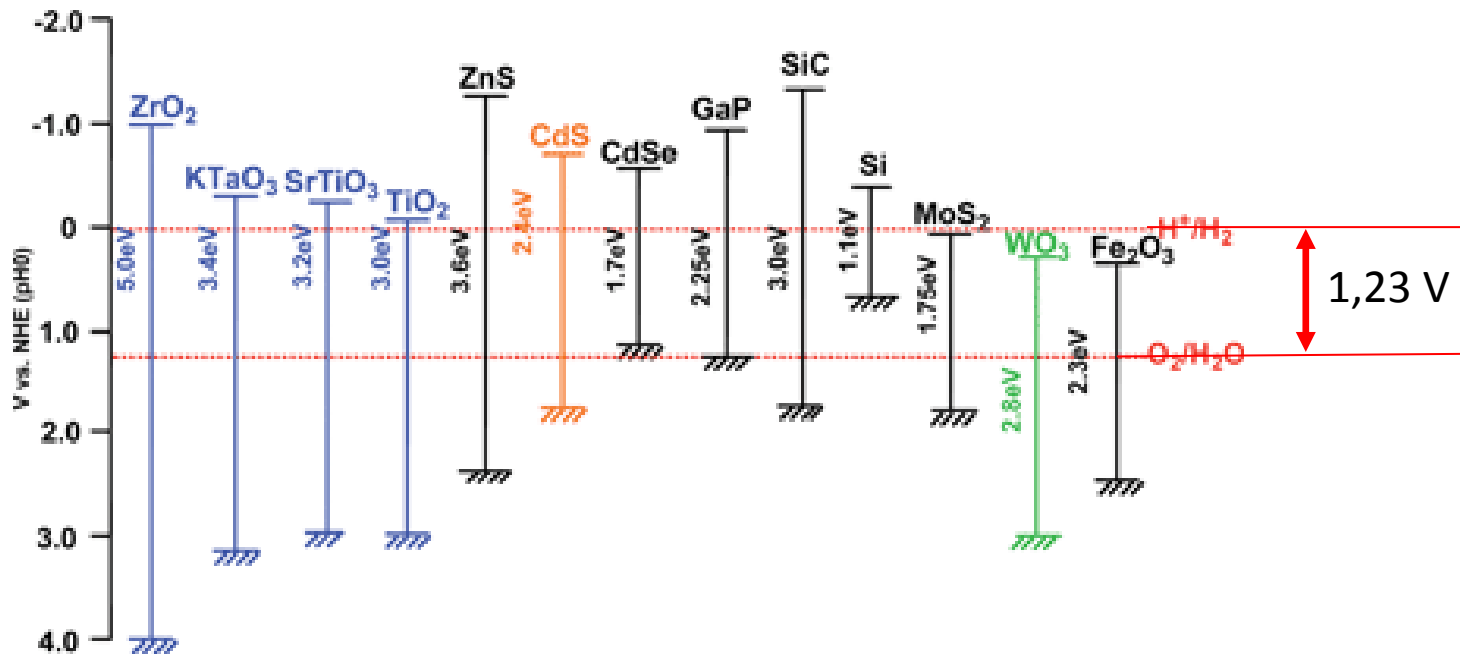


# Coupled elementary processes in a PEC

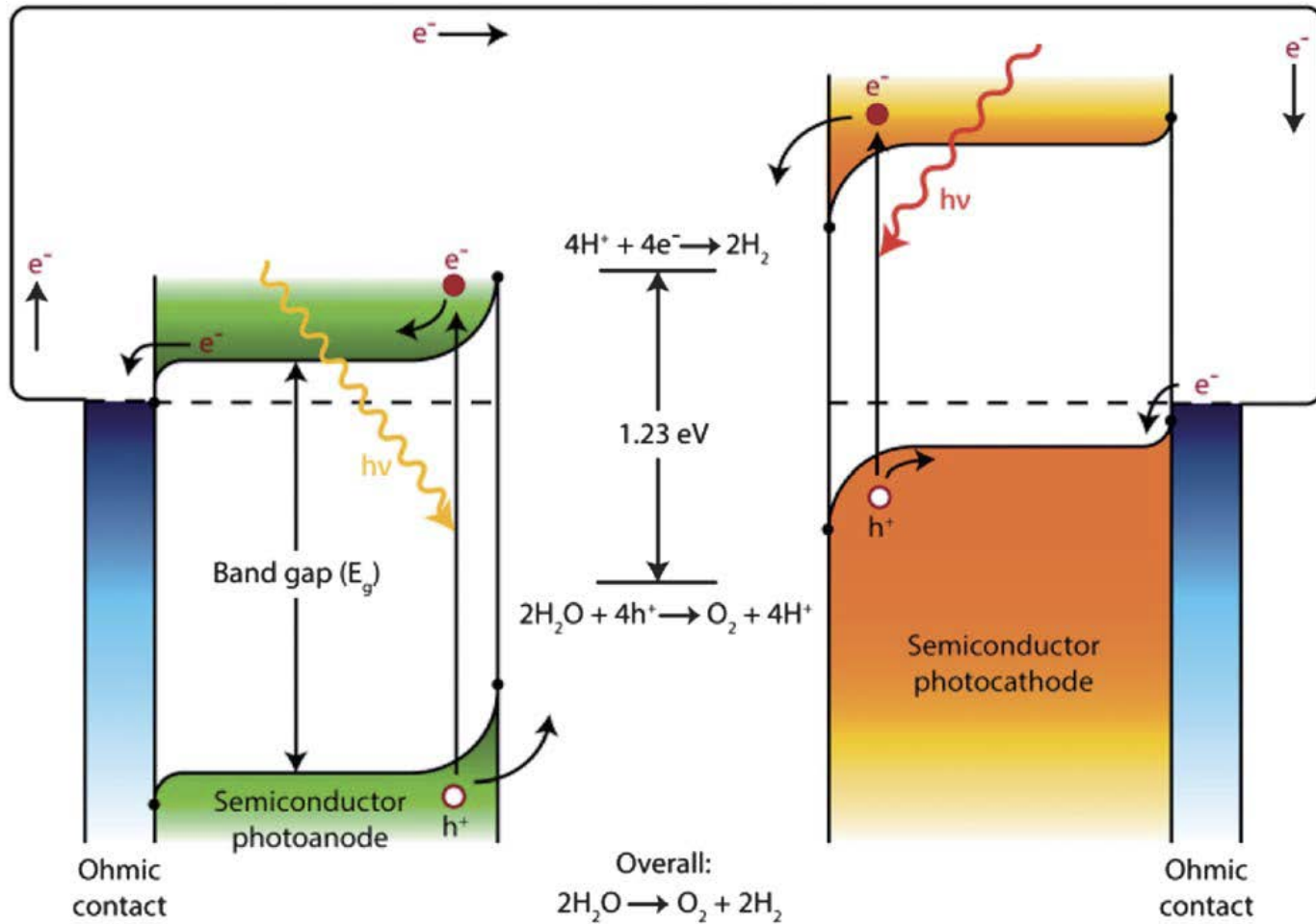


# Energetics

To drive water splitting:  $U \gg 1,23 \text{ V}$   $\longrightarrow$  Dual junction cells



# Band Diagram



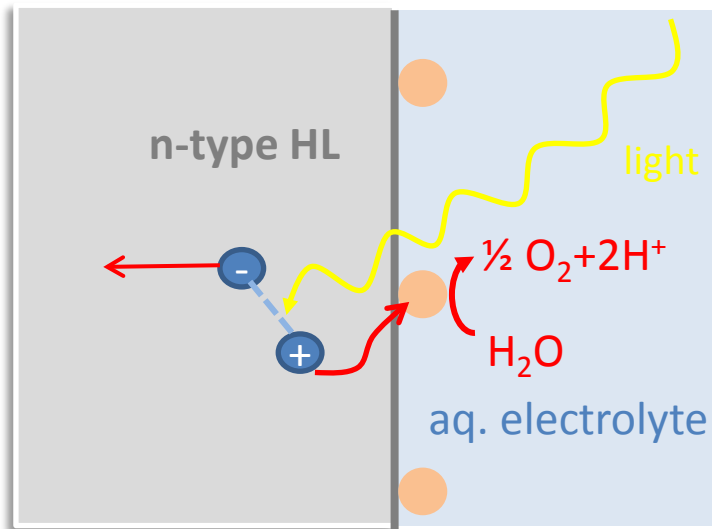
# Alternative solar fuels from CO<sub>2</sub> reduction

| Reaction                                                                   | $E_0$ [V] |
|----------------------------------------------------------------------------|-----------|
| $O_2(g) + 4 H^+(aq) + 4 e^- \rightleftharpoons 2 H_2O(l)$                  | 1.23      |
| $2 H^+(aq) + 2 e^- \rightleftharpoons H_2(g)$                              | 0.00      |
| $CO_2(g) + 2 H^+(aq) + 2 e^- \rightleftharpoons CO(g) + H_2O(l)$           | -0.11     |
| $CO_2(g) + 8 H^+(aq) + 8 e^- \rightleftharpoons CH_4(g) + 2 H_2O(l)$       | 0.17      |
| $CO_2(g) + 2 H^+(aq) + 2 e^- \rightleftharpoons HCOOH(l)$                  | -0.25     |
| $2 CO_2(g) + 12 H^+(aq) + 12 e^- \rightleftharpoons C_2H_4(g) + 4 H_2O(l)$ | 0.08      |

- Suppression of H<sub>2</sub> evolution
- Selectivity

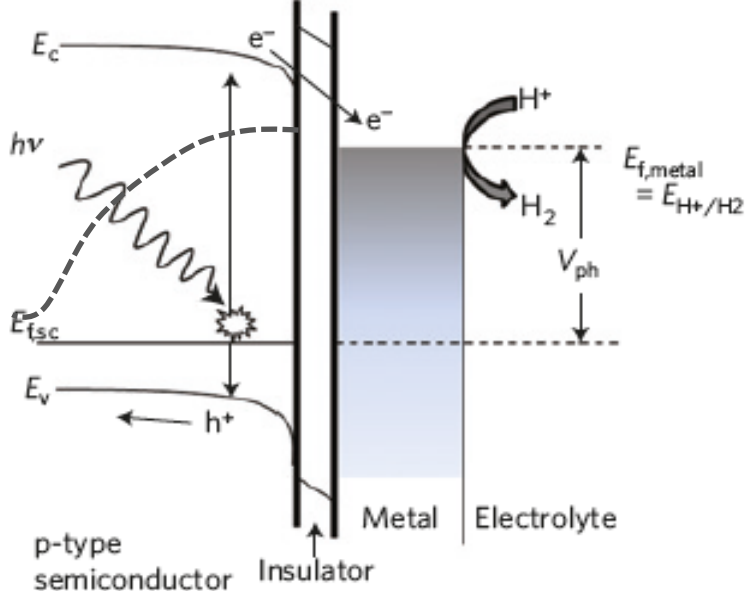
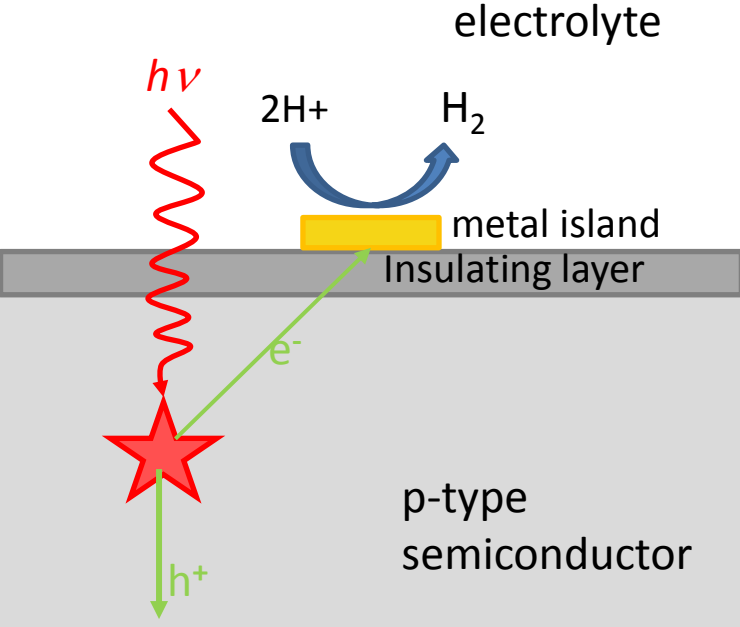
# Optimization of the Semiconductor- Electrolyte Interface

## Requirements for high efficiency



- High catalytic activity of the interface
- Sufficient passivation of the surface (low rate of electron-hole recombination)
- Adjustment of energy levels of catalyst and reaction
- No further parasitic energy levels in the electrolyte

# Metal/Insulator/Semiconductor Photoelectrode





# *Cu particle decorated SiO<sub>2</sub>/p-Si electrodes*

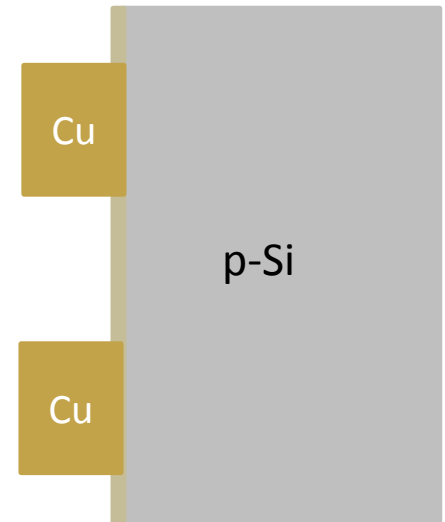
## Combining two well studied subsystems

### Advantages Si:

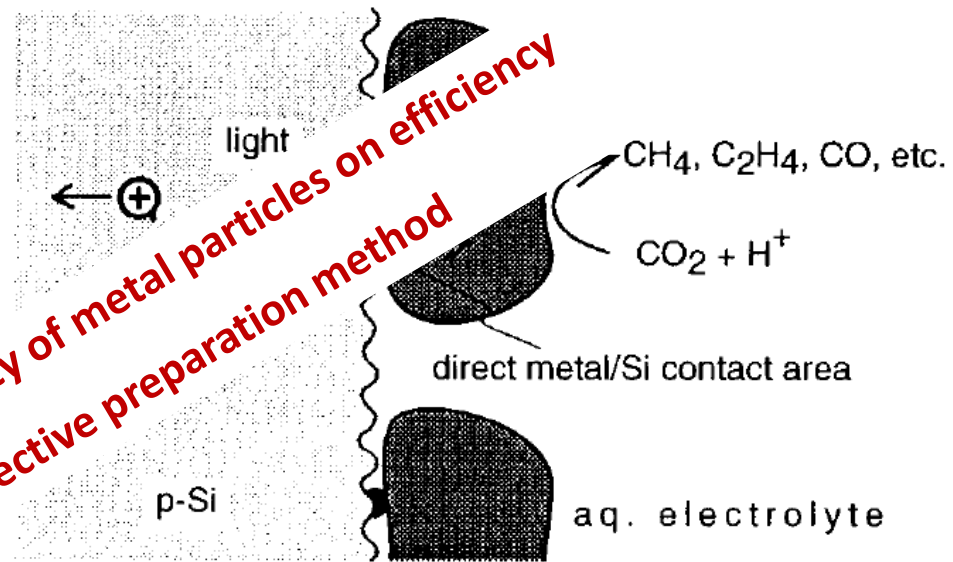
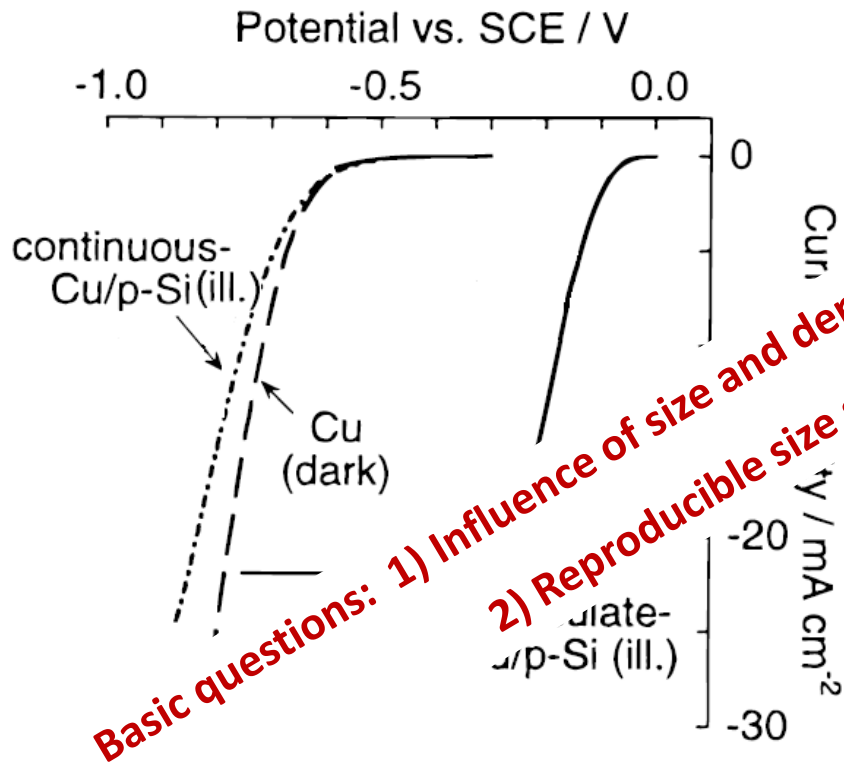
- abundant, sustainable
- band gap well matched to solar spectrum
- widely used in photovoltaics → technology available, low cost material
- SiO<sub>2</sub> insulating layer stable over wide pH and potential ranges

### Advantages Cu:

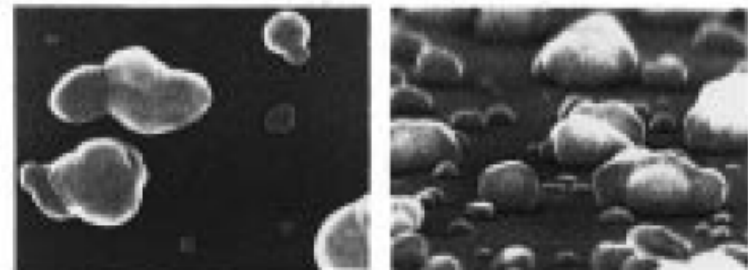
- Electrochemically active for reduction of CO<sub>2</sub> to methane or ethylene
- Electrochemistry well studied



# Current understanding: CO<sub>2</sub> reduction on p-Si/Cu photoelectrodes

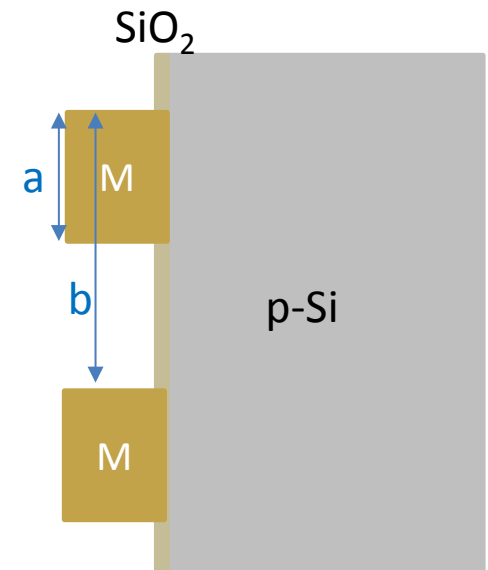


SEM image:



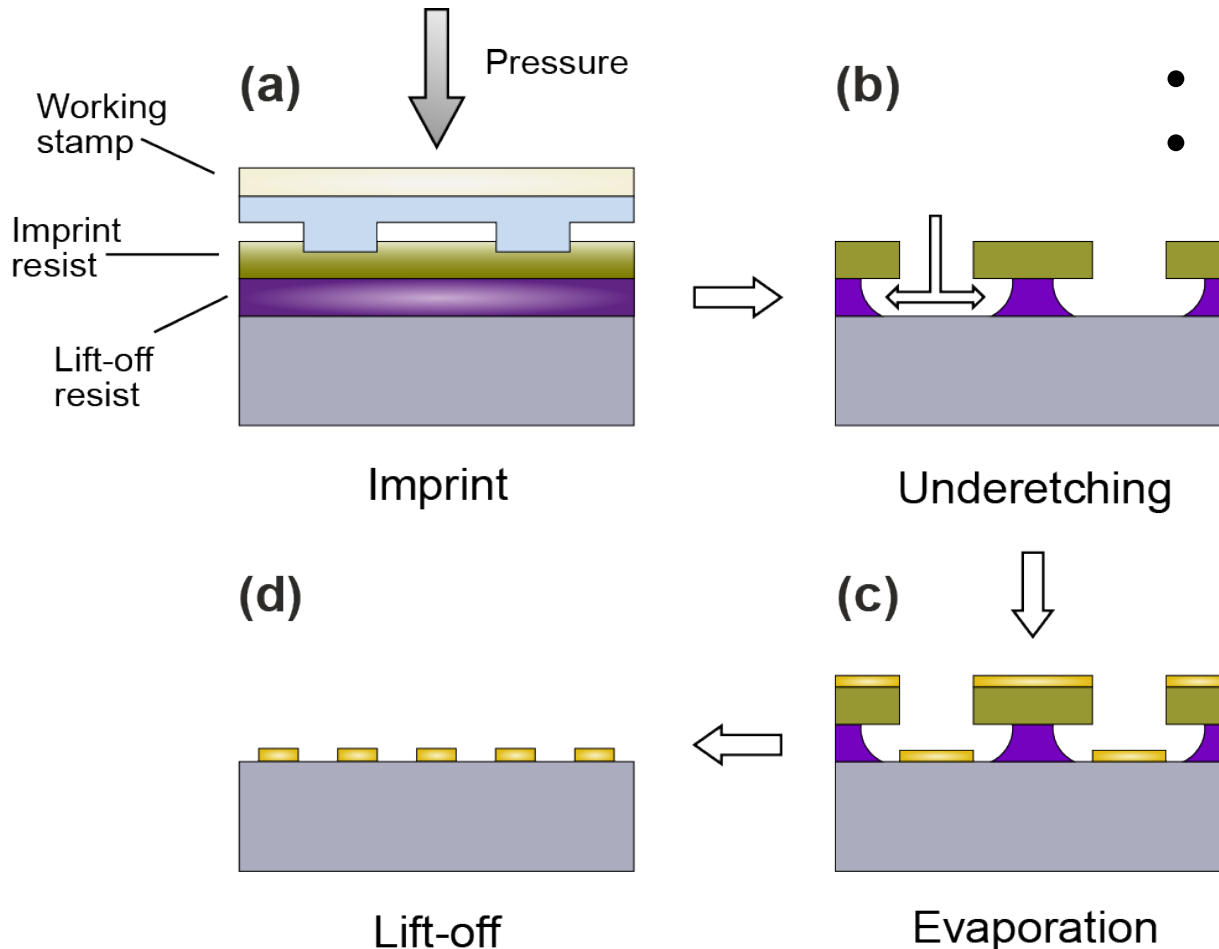
# ***Ideal design of the Si/SiO<sub>x</sub>/metal interface?***

- Ideal structure size  $a$
- Ratio of structure size  $a$  to pitch size  $b$
- Influence on structure size on energetics and catalytic activity



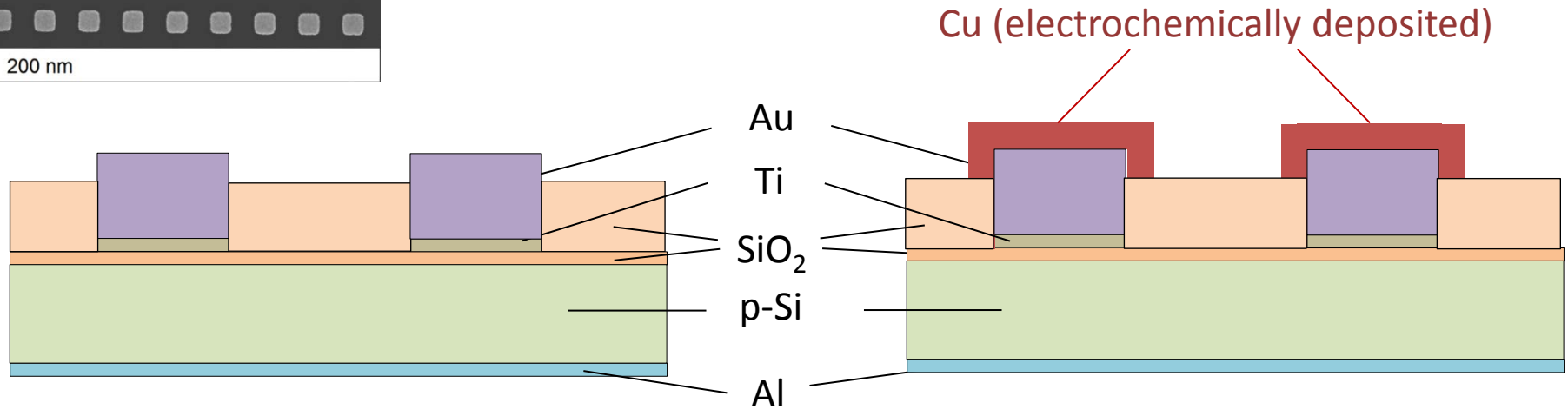
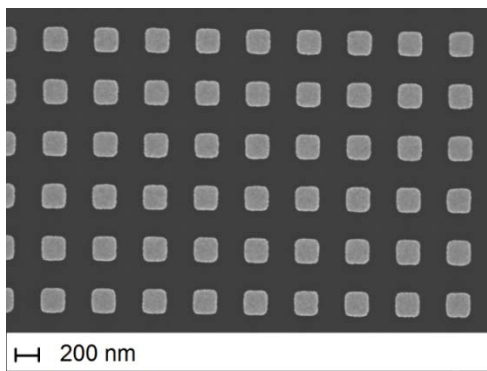
# Lift-off nanoimprint lithography

large-scale deposition of metal structures (cm<sup>2</sup> range)



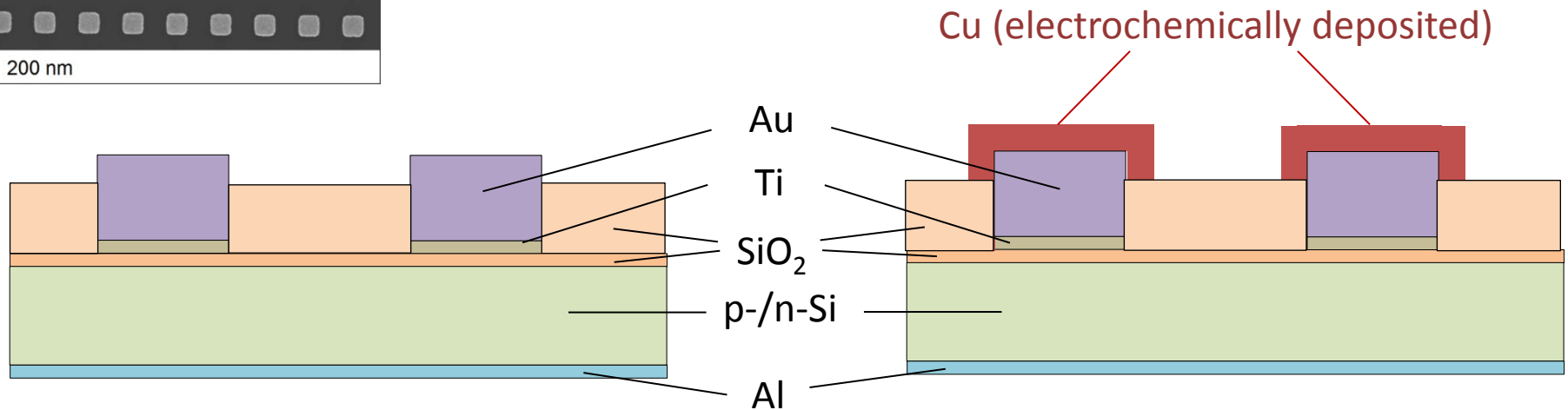
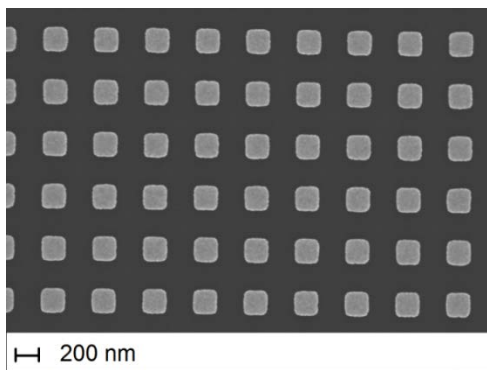
- Feature size down to 45nm
- Feature size tuning

# Final Electrode Design



- SiO<sub>2</sub> (~ 1nm): diffusion barrier layer
- Ti (~ 0.5 nm): adhesion layer
- Thermal oxide (~ 15nm): passivating layer
- Au (Cu): electrochemically active area

# Final Electrode Design

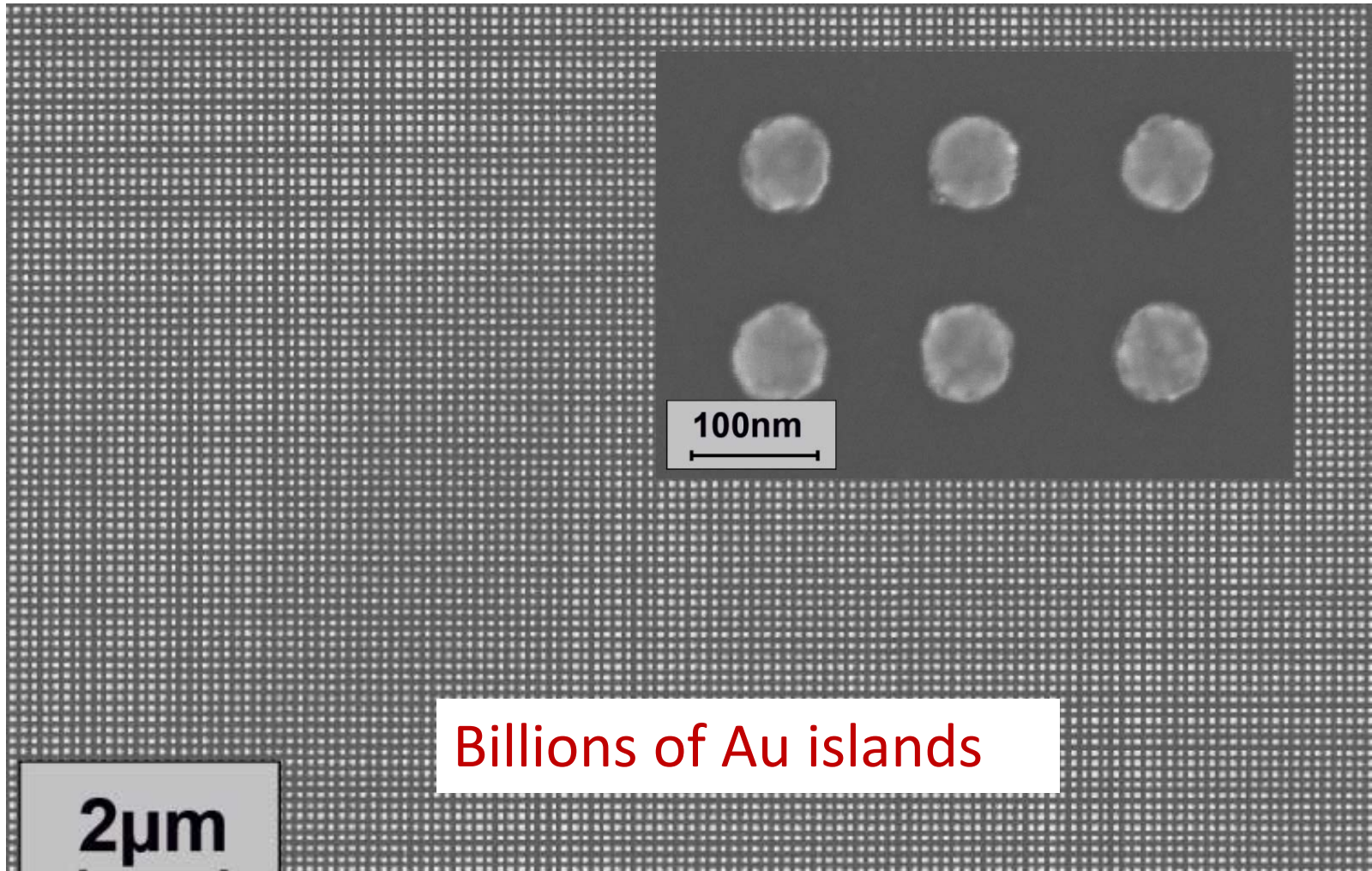


- SiO<sub>2</sub> (~ 1nm): diffusion barrier layer
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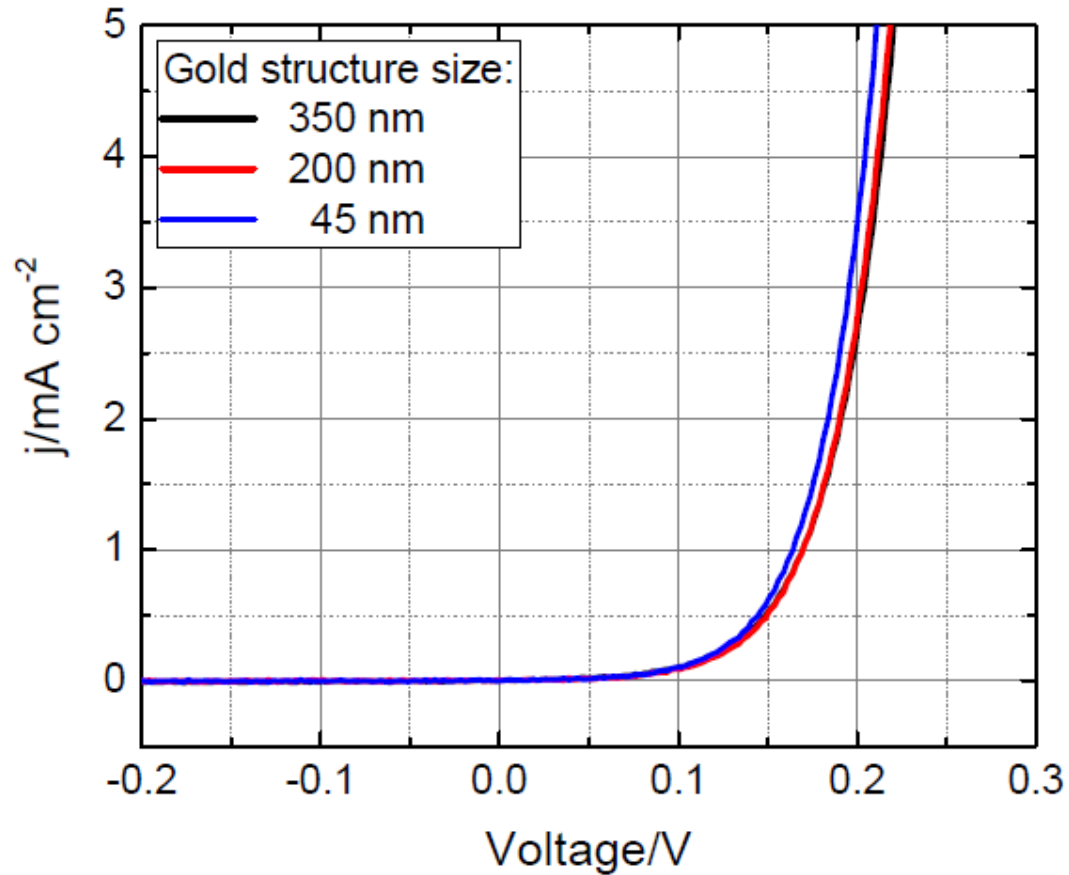


- Electrochemically stable
- Homogeneously distributed and well defined
- Tunable in size (down to 45 nm diameter possible)

# Lift-off nanoimprint lithography



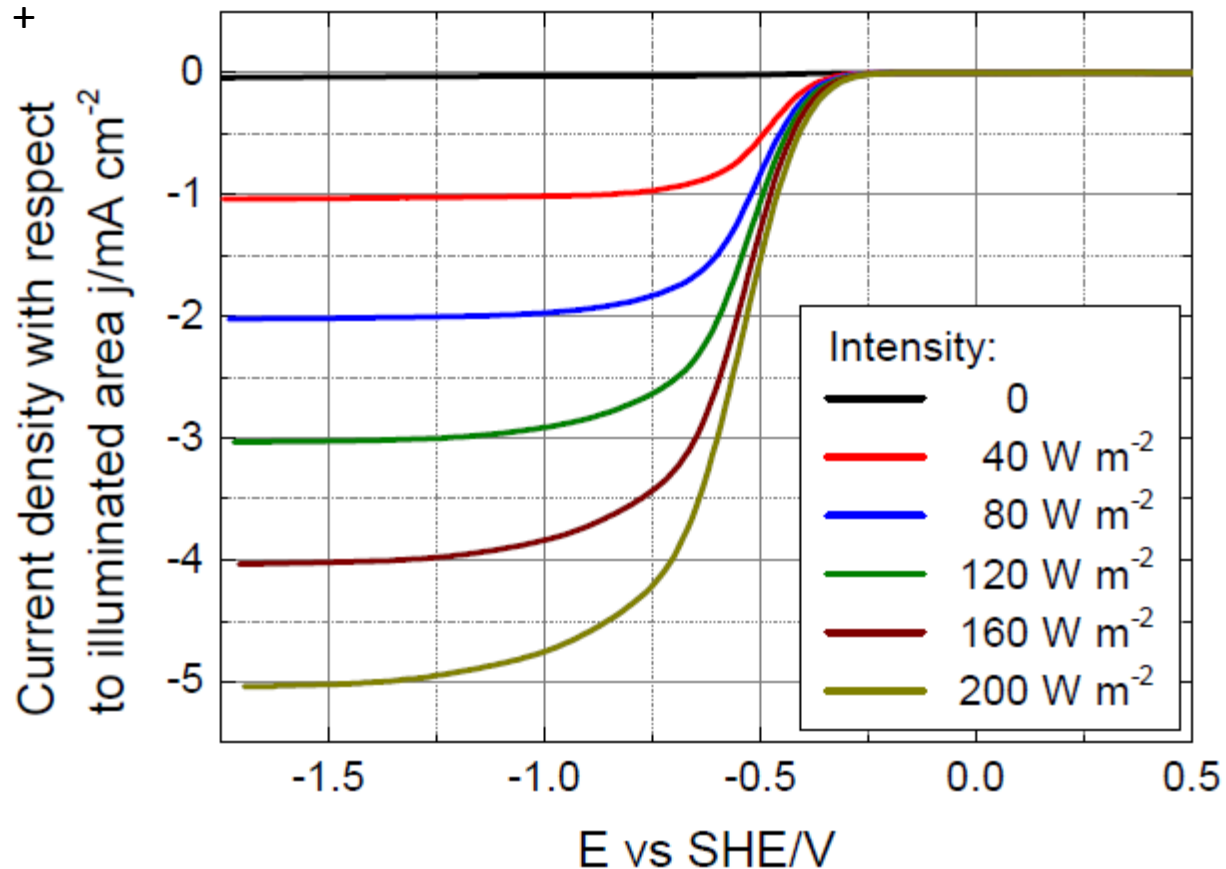
# „Dry‘ $I$ - $U$ curves of p-Si/SiO<sub>2</sub>/Ti/Au nanostructure



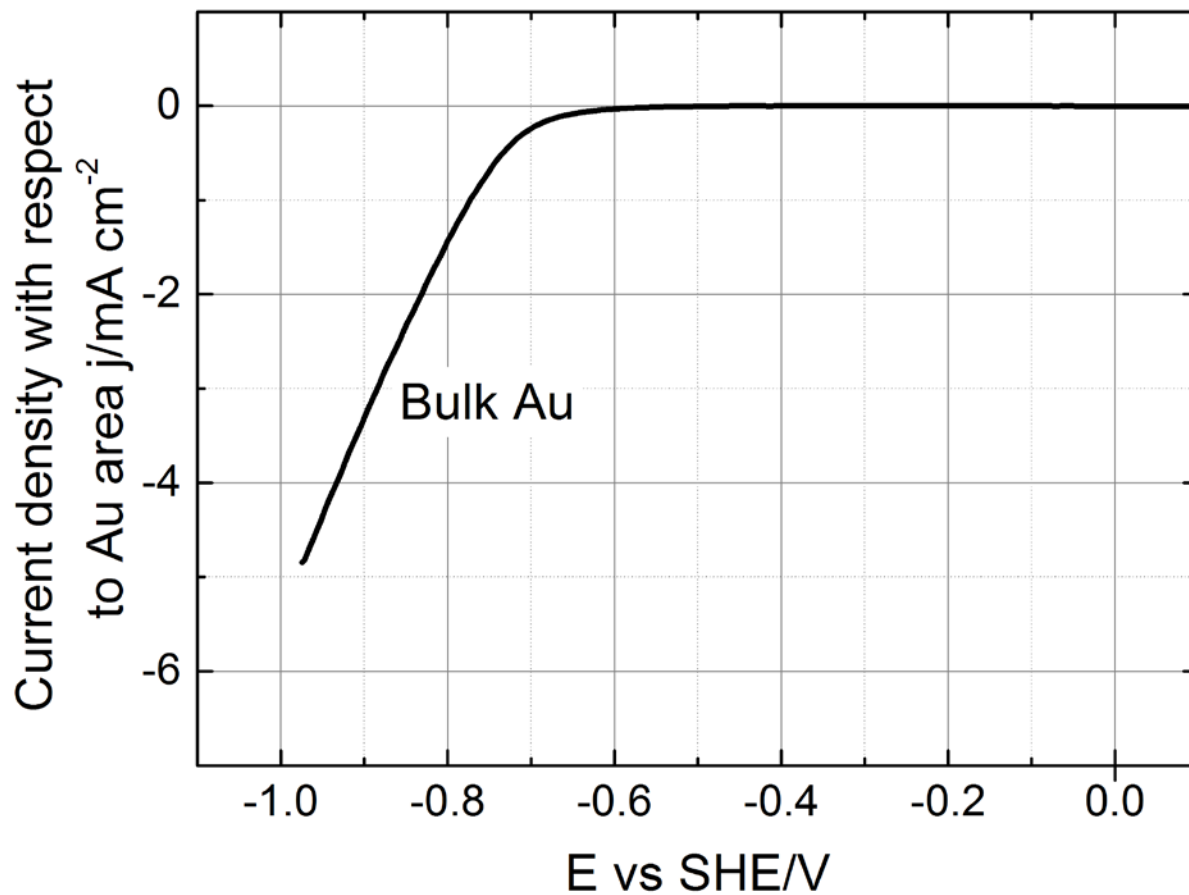


# *I-U* curves of the Au/Ti/SiO<sub>2</sub>/p-Si/electrolyte interface for different illumination intensities

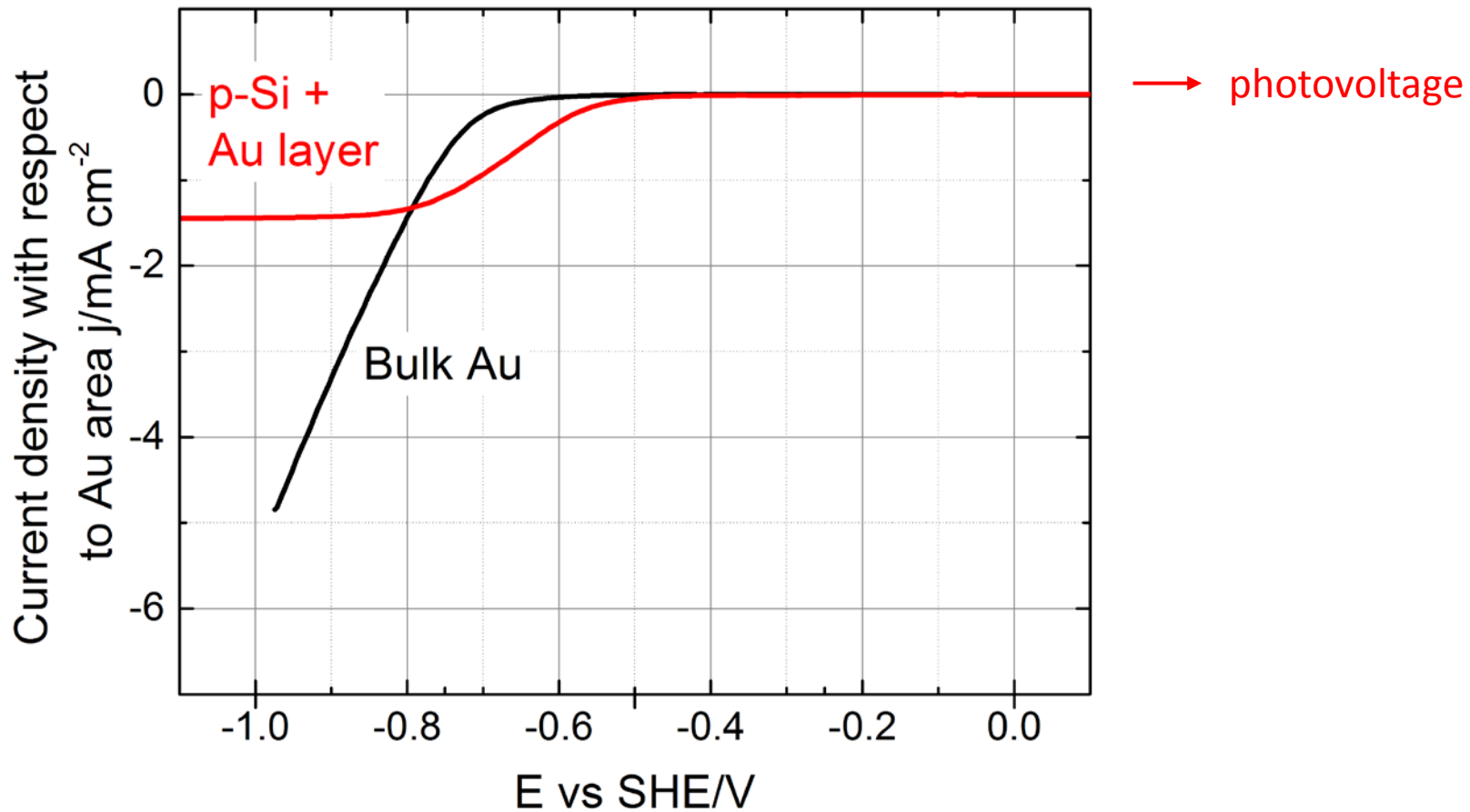
- WE: 200 nm Au squares
- Electrolyte: 75 mM K<sub>2</sub>CO<sub>3</sub> + 100 mM H<sub>3</sub>PO<sub>4</sub> purged with CO<sub>2</sub>



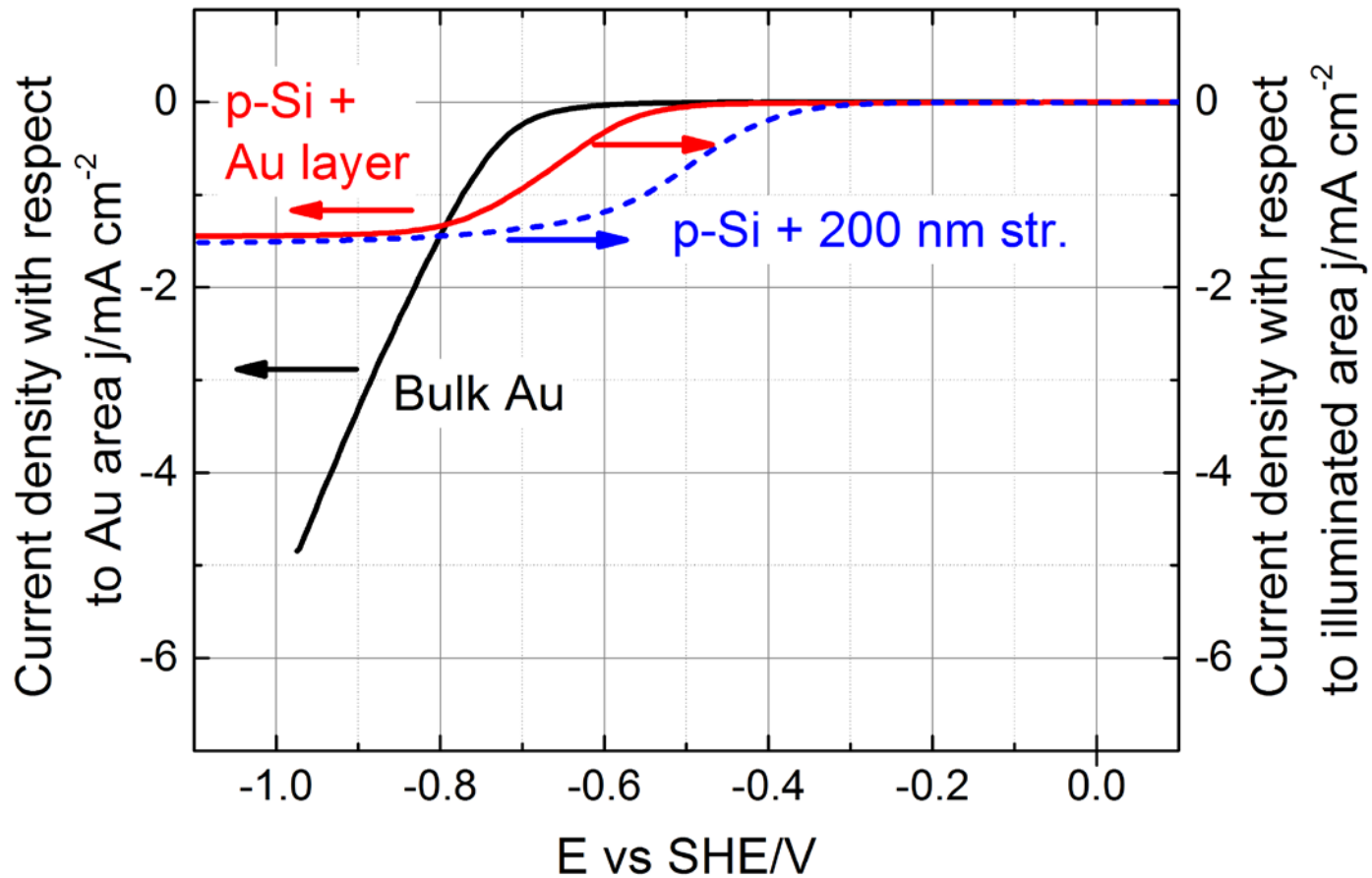
# ***Comparison of continuous Au layer and nanostructured Au on p-Si/SiO<sub>2</sub>***



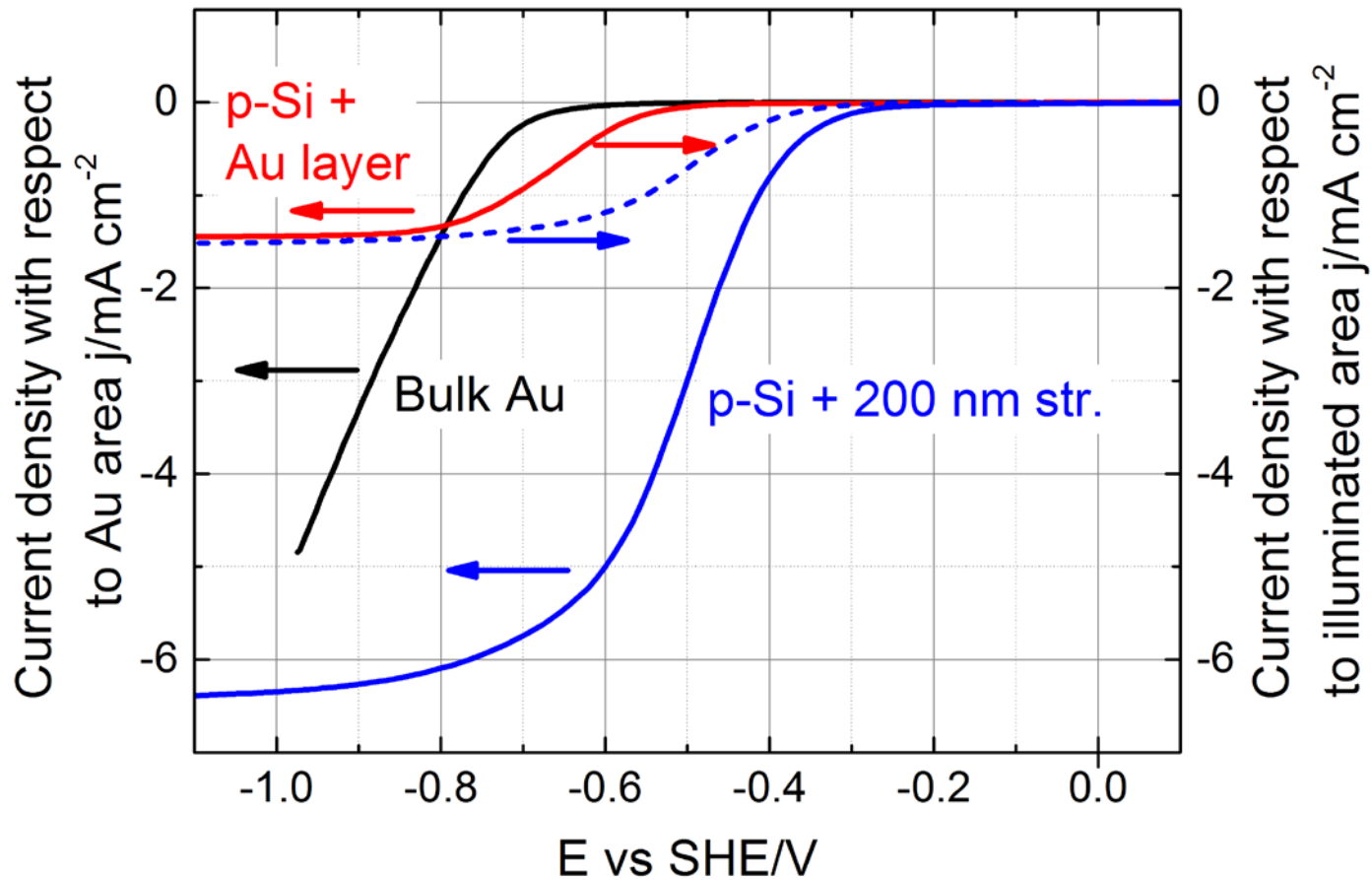
# Comparison of bulk Au, continuous Au layer and nanostructured Au on SiO<sub>2</sub>/p-Si



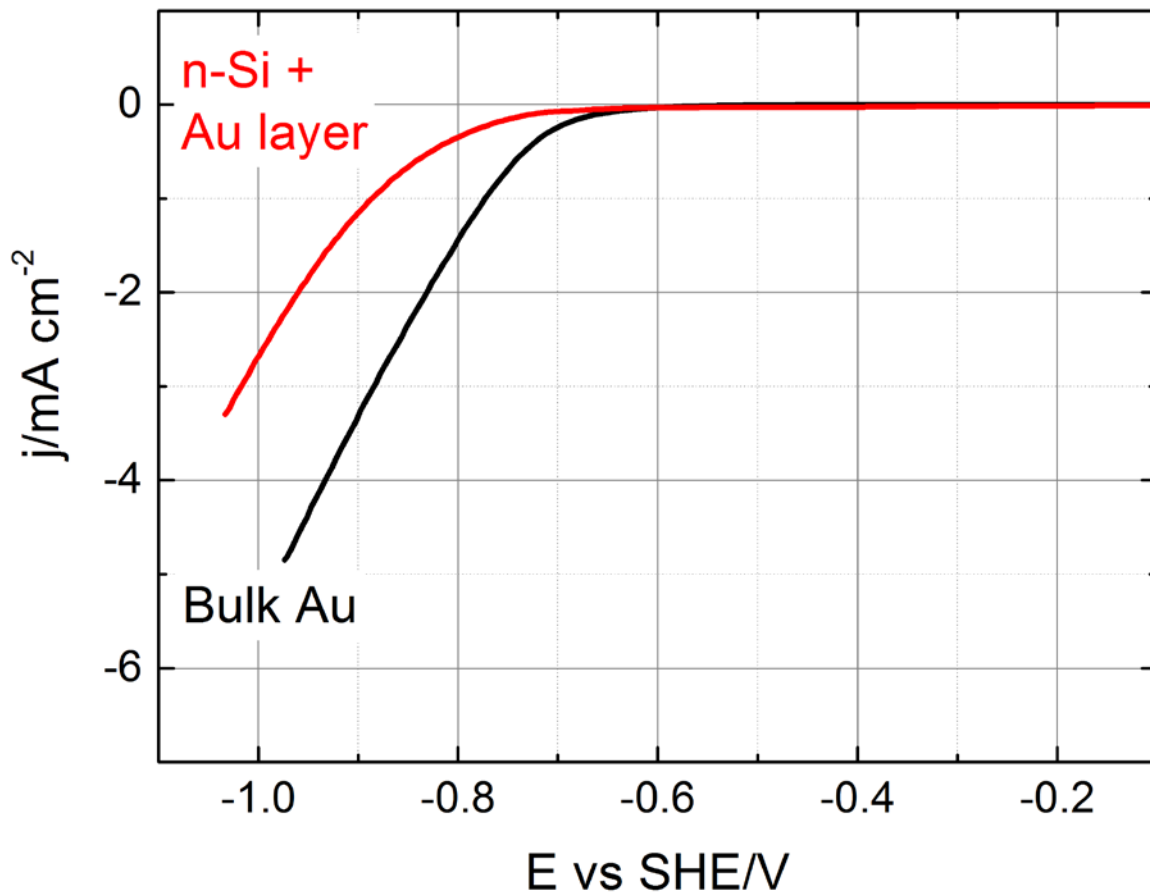
# Comparison of bulk Au, continuous Au layer and nanostructured Au on p-Si/SiO<sub>2</sub>



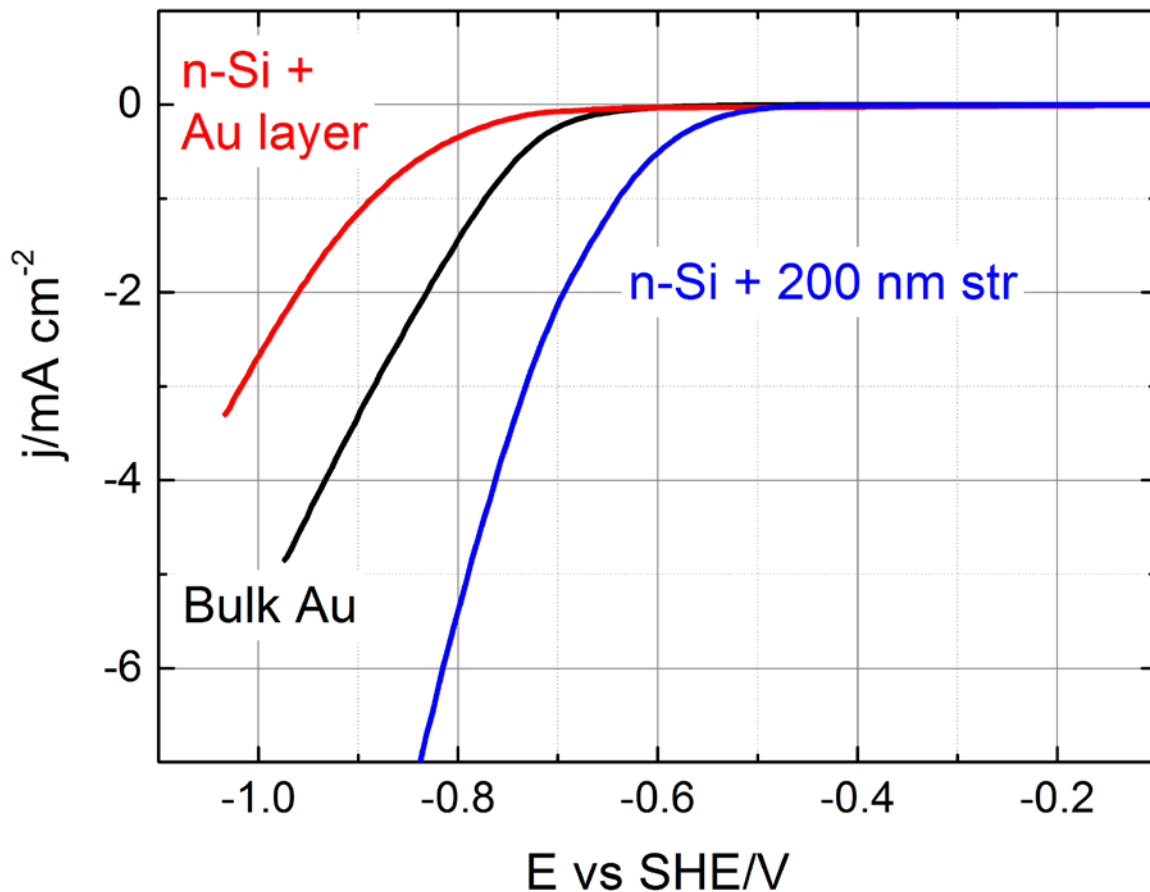
# Comparison of bulk Au, continuous Au layer and nanostructured Au on p-Si/SiO<sub>2</sub>



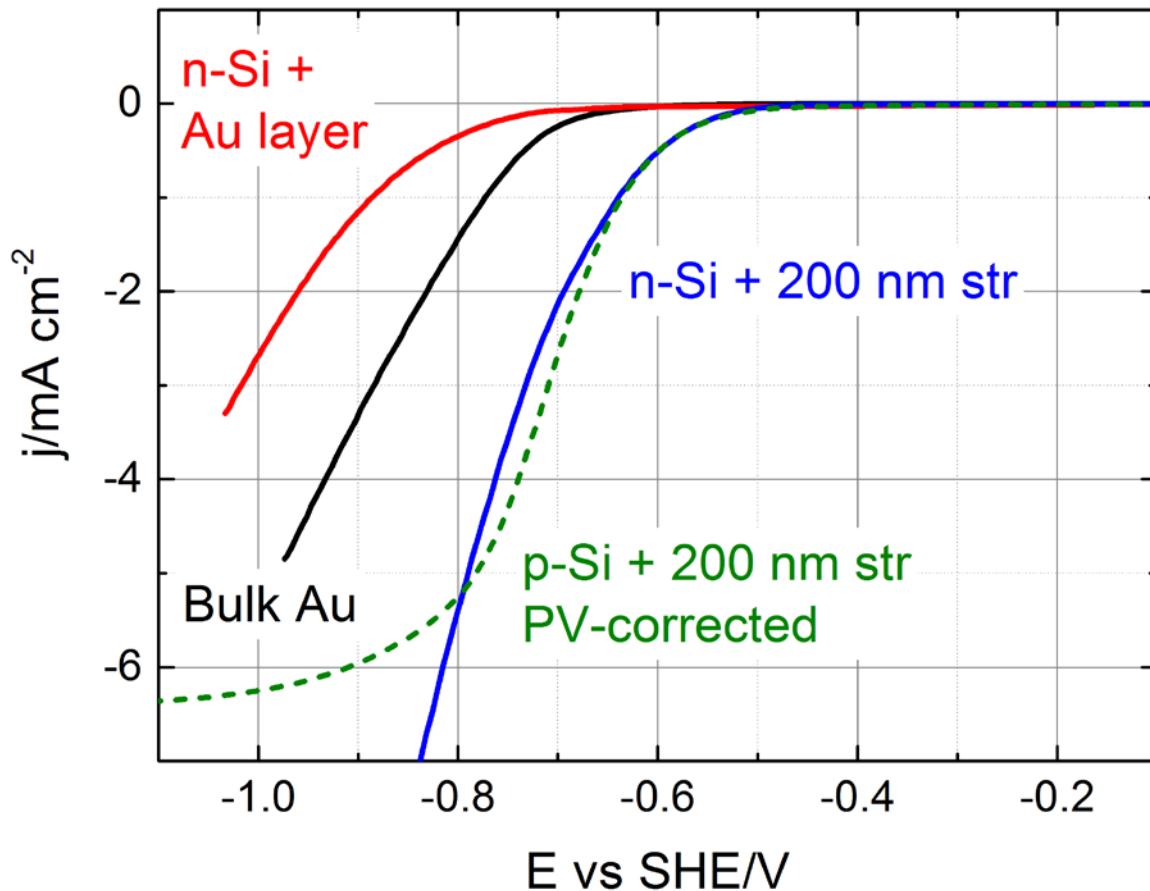
# ***Comparison of bulk Au and n-Si and p-Si based MIS electrodes***



# Comparison of bulk Au and n-Si and p-Si based MIS electrodes

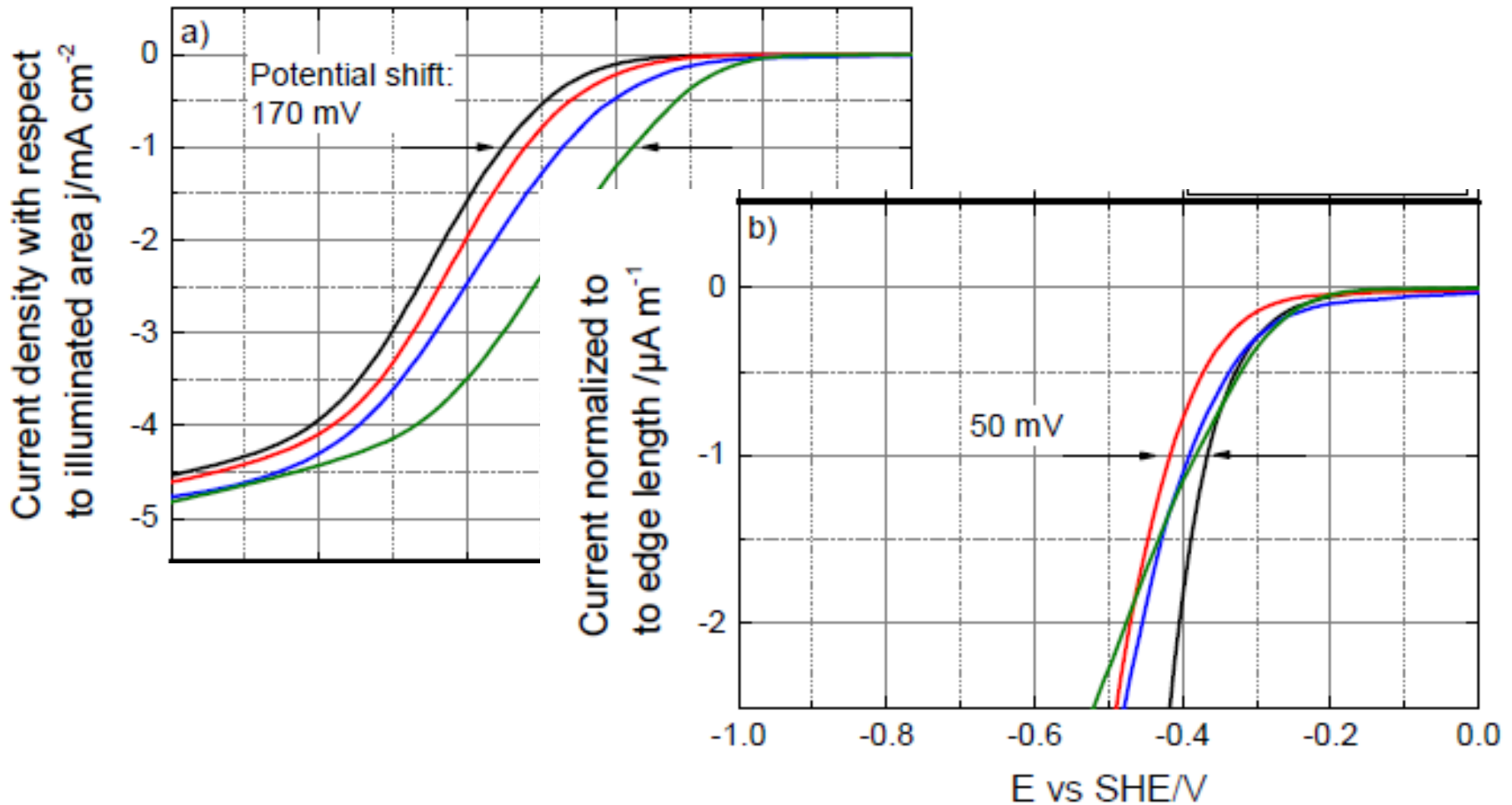


# Comparison of bulk Au and n-Si and p-Si based MIS electrodes

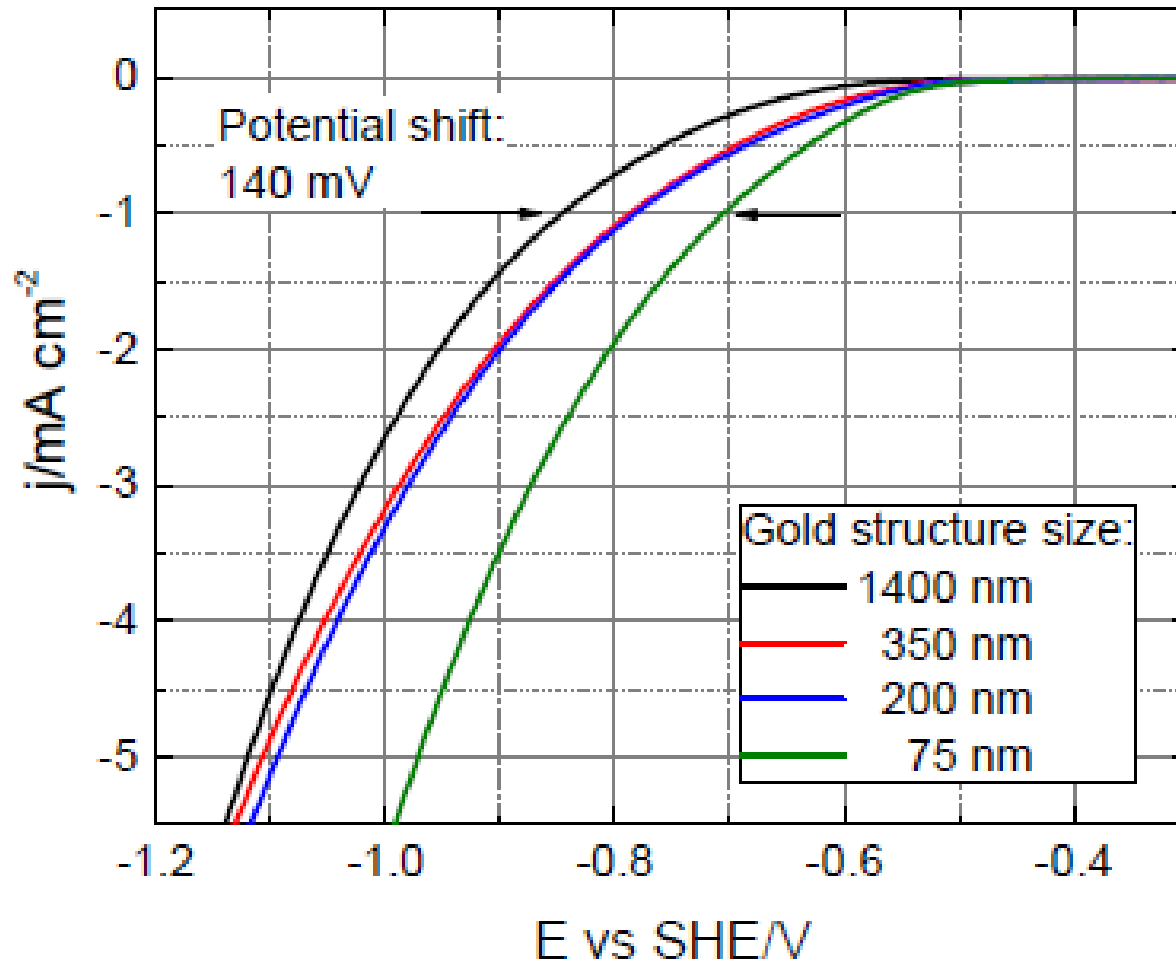




# p-Si/SiO<sub>2</sub>/Au: Influence of structure size

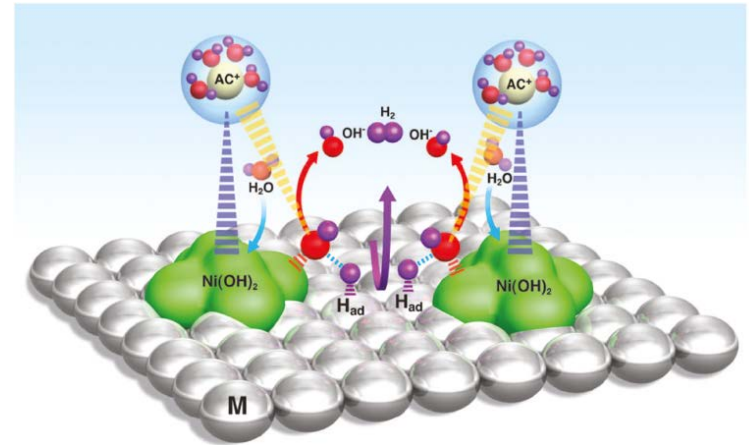
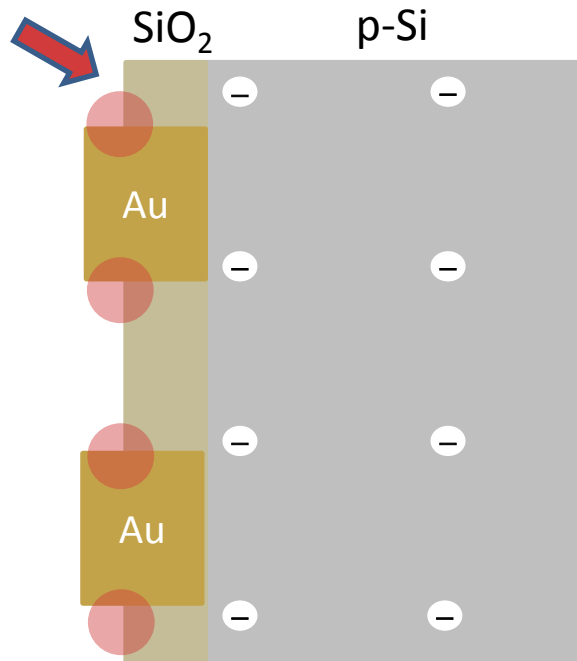


# Degenerately doped $p^{++}$ -Si/SiO<sub>2</sub>/Au: Influence of structure size

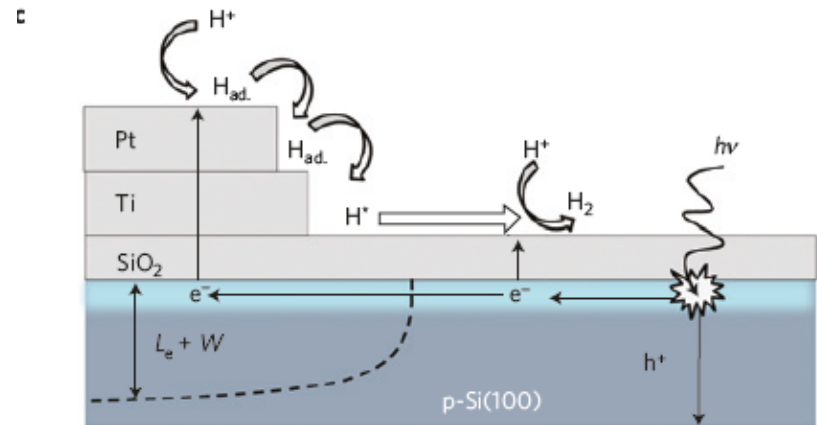


# Enhanced reactivity at edges → bifunctional mechanism?

Chemical modification of reactivity along the rime



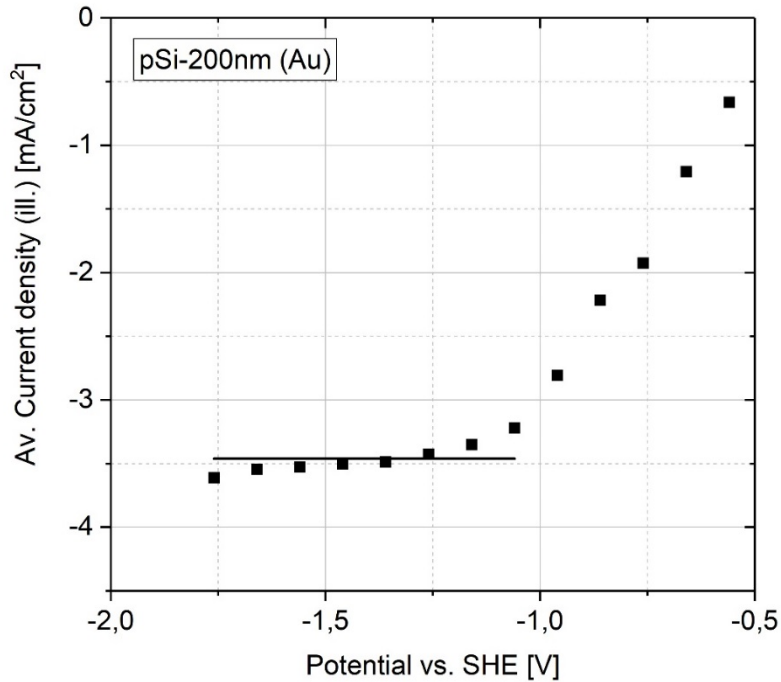
Subbaraman et al., Science (2011)



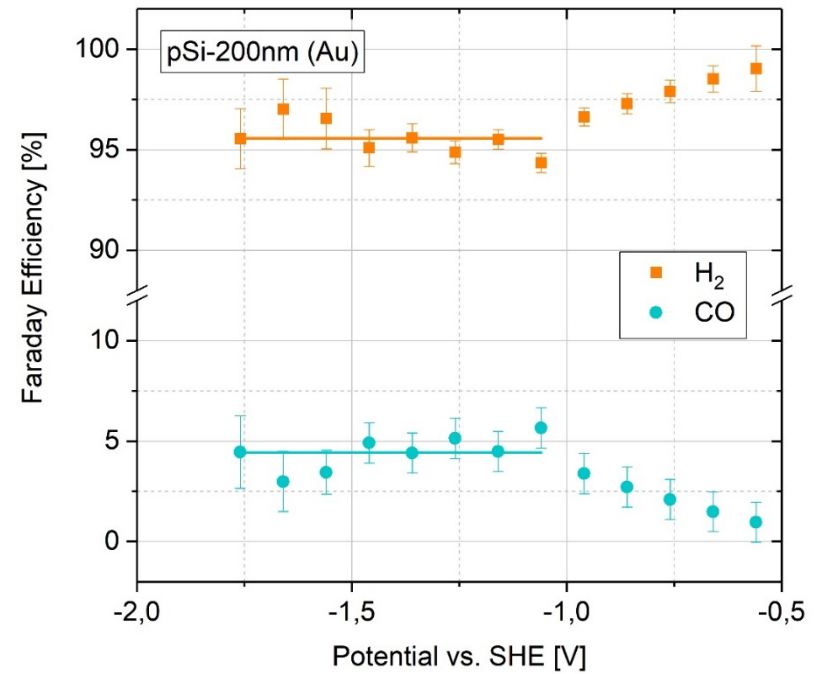
Esposito et al., Nature Materials (2013)

# Electrolysis measurements: Au arrays

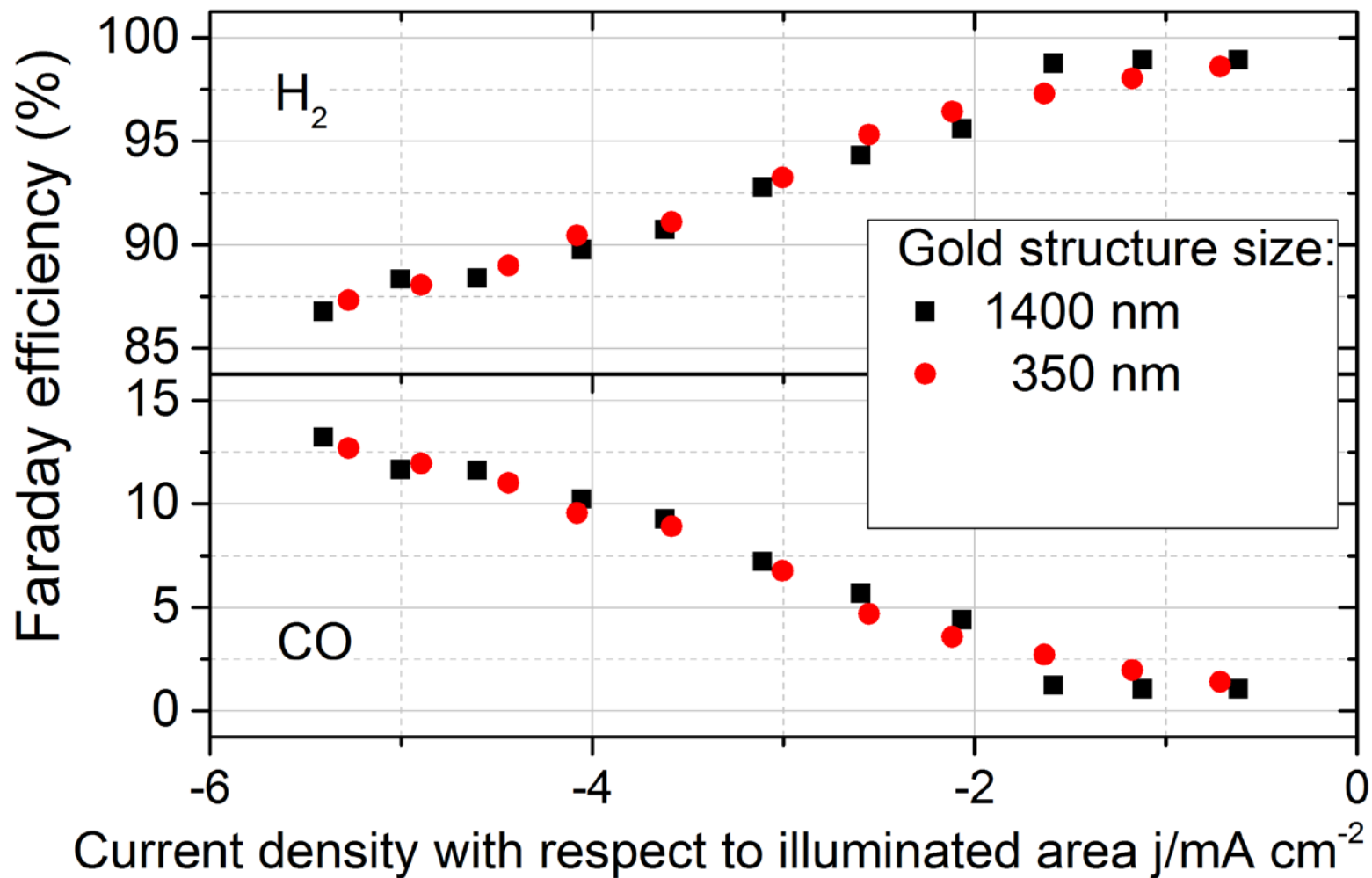
## Av. Current vs. Potential



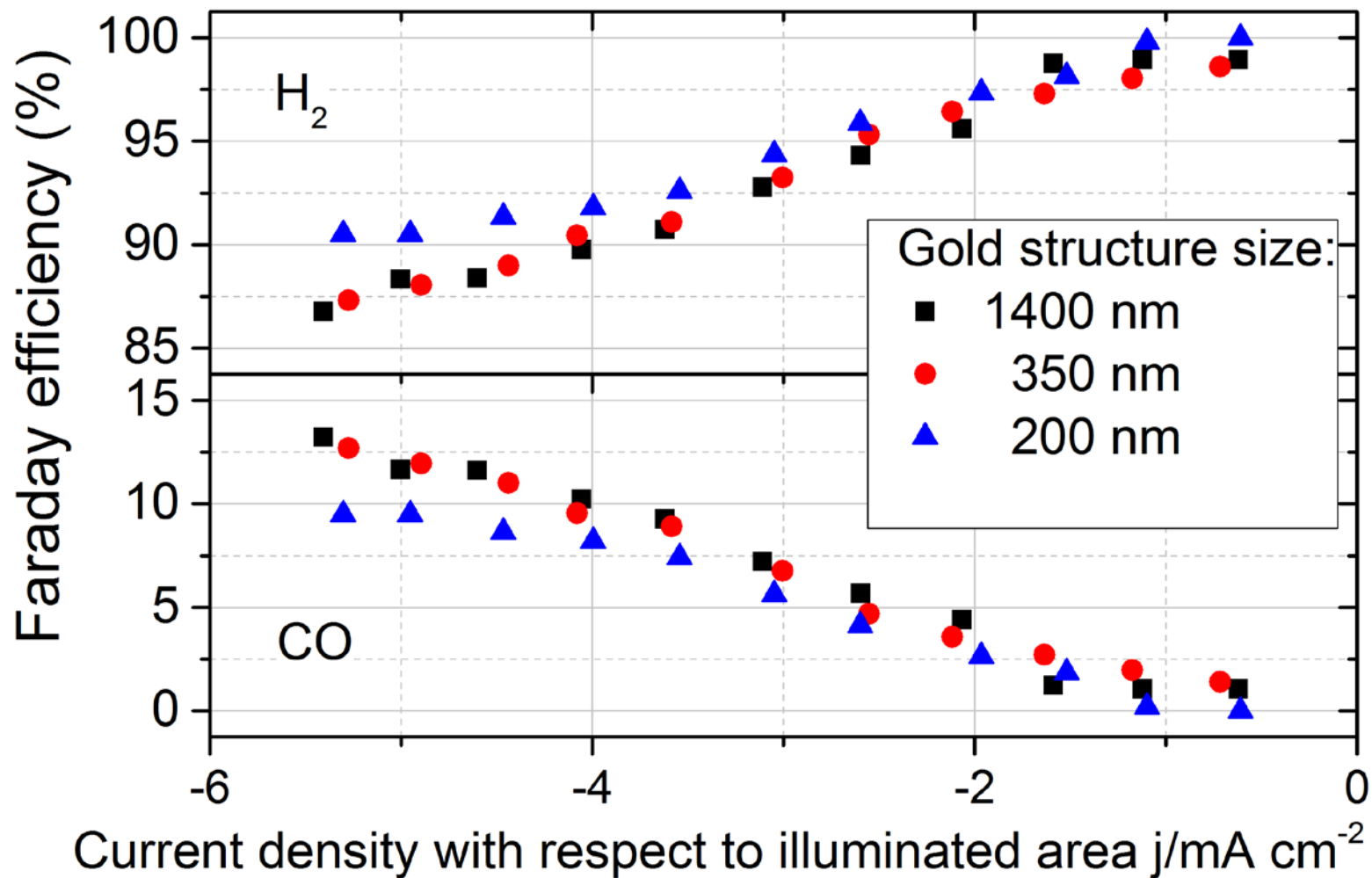
## Product distribution vs. Potential



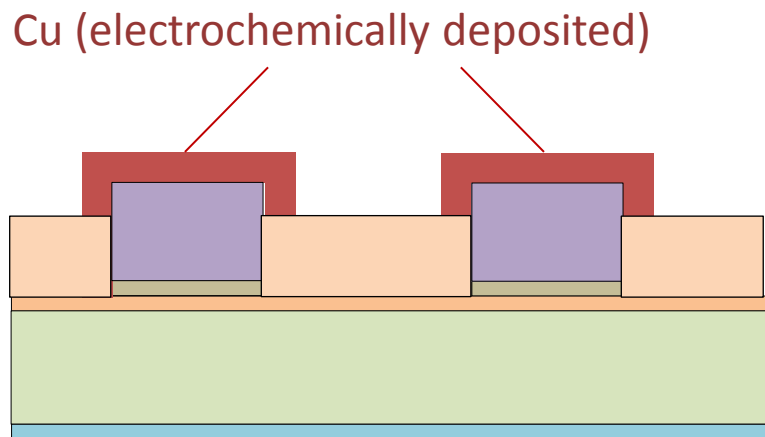
# Product distribution



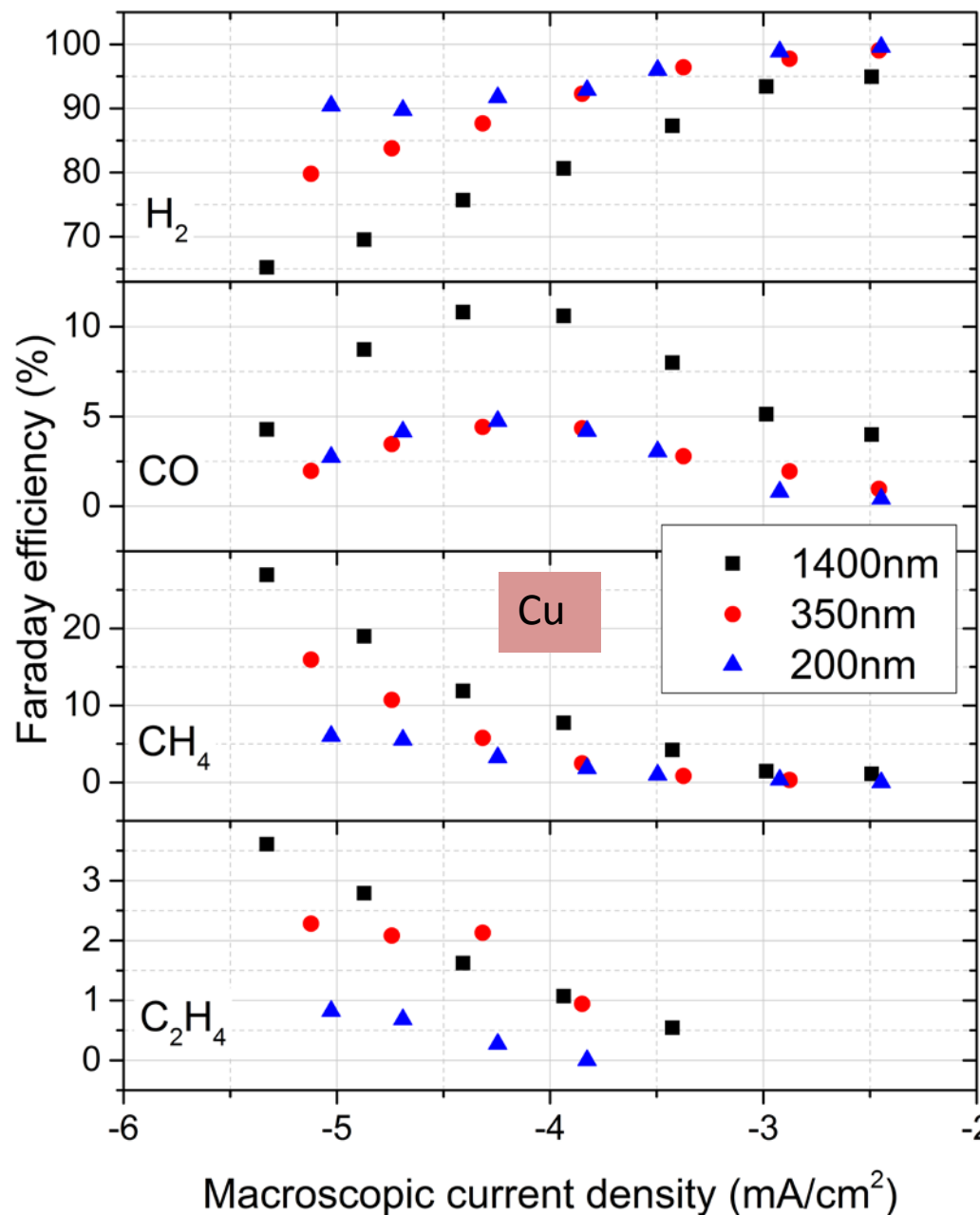
# Product distribution



# Cu-coated Au arrays

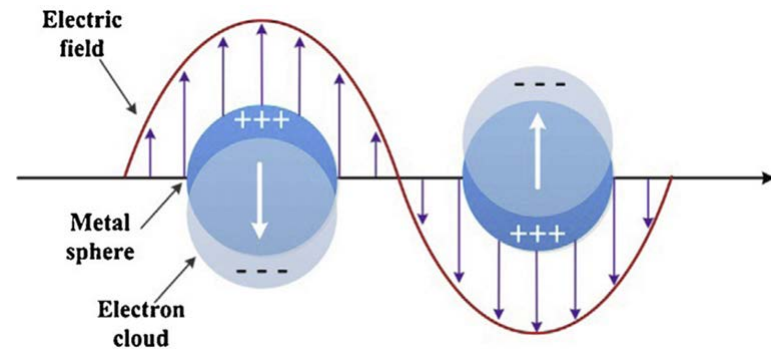


With decreasing size: reduced hydrocarbon efficiency



# Localized surface plasmon (LSP)

***LSP: Surface plasmon confined in a nanoparticle of size comparable to or smaller than the wavelength of light used to excite the plasmon.***

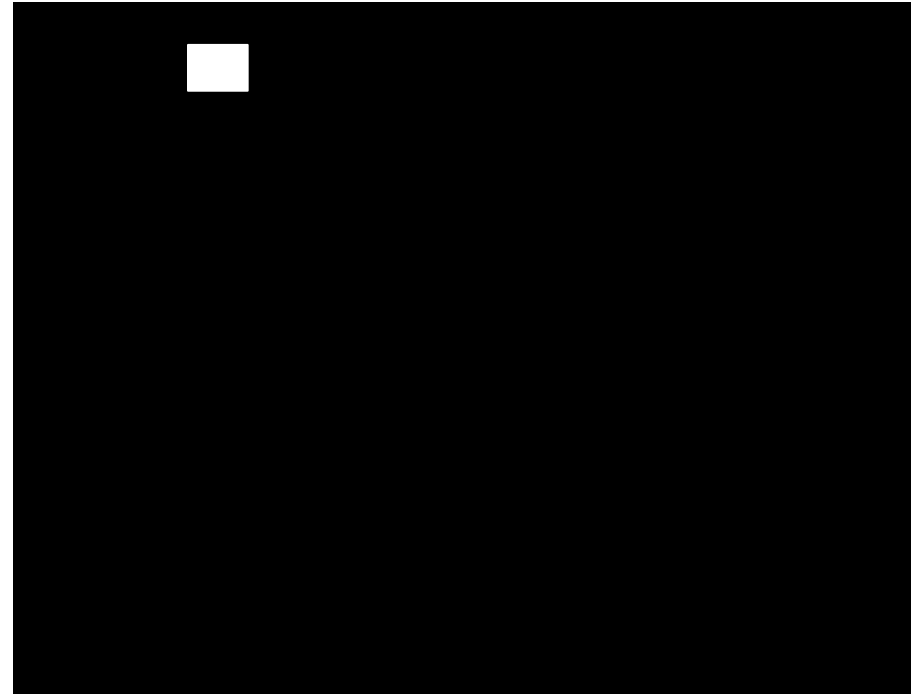
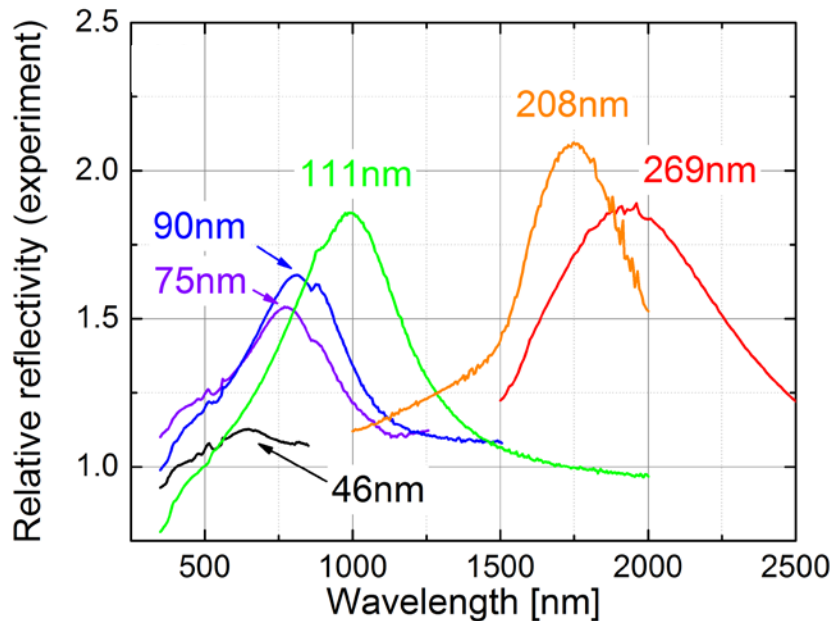


Sun, Chen and Lin, DOI: 10.5772/64380

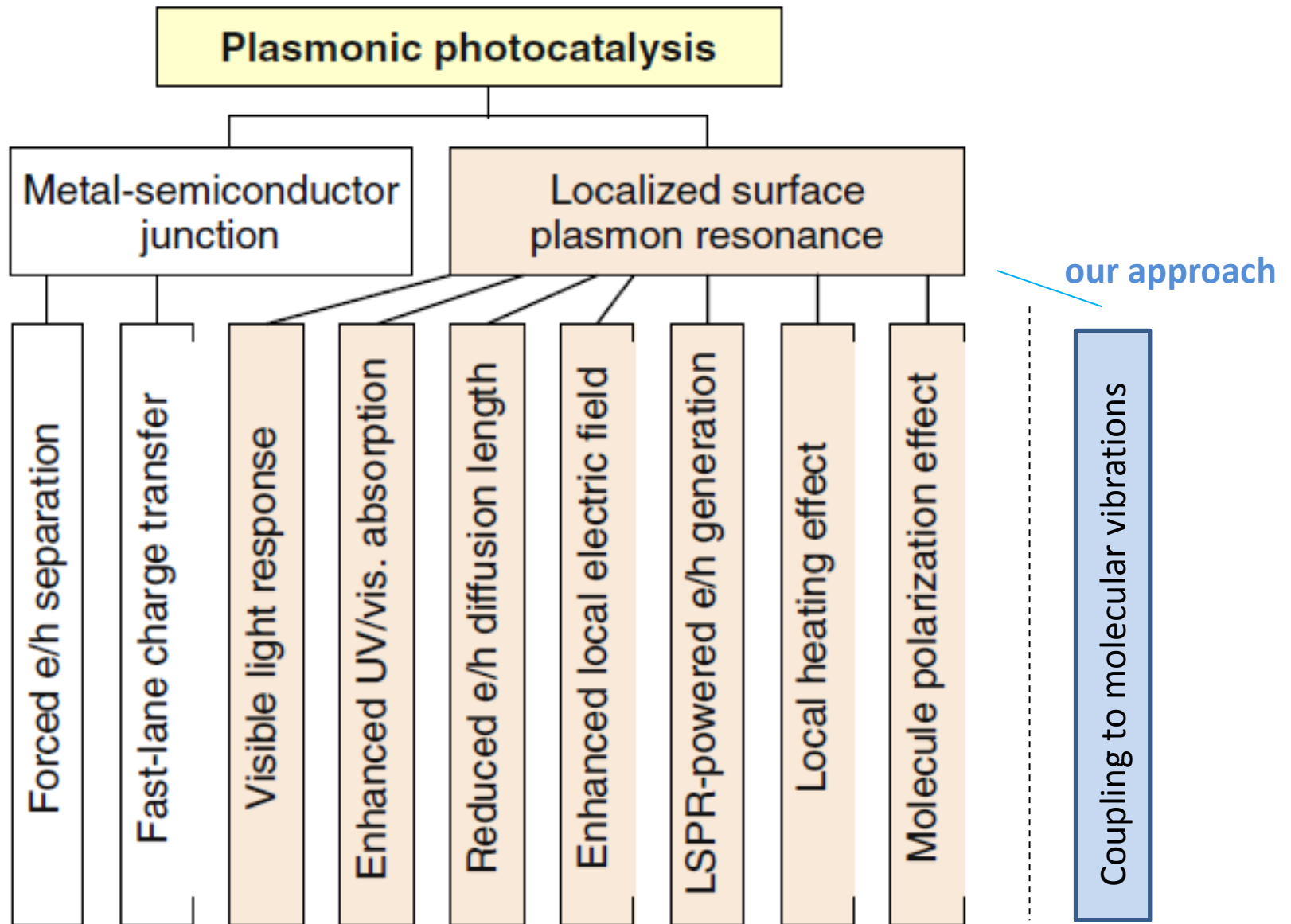
- strongly enhanced electric fields near the particle's surface
- Strong absorption around the plasmon resonance frequency



# Tunable plasmonic resonances



- UV-Vis spectra and simulations of differently sized gold structures in good agreement

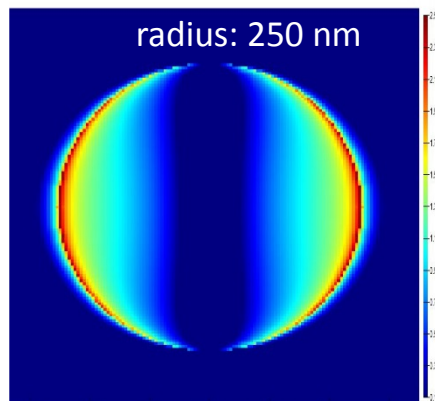


# ***Coupling of LSPs to molecular vibrations***

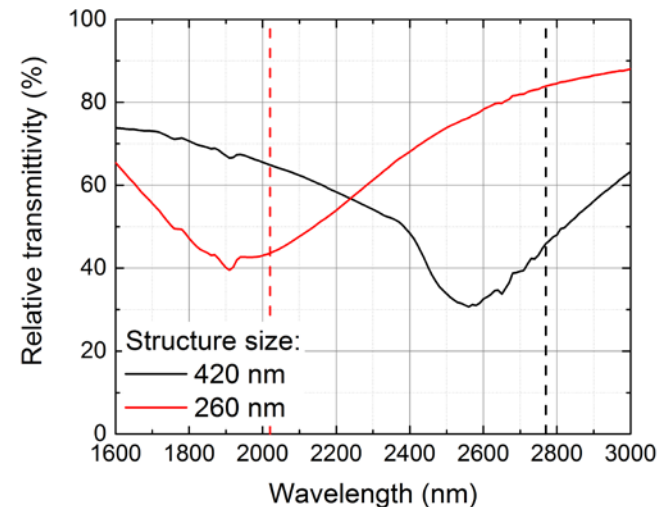
- Reduction of activation energy of a reaction → Plasmonic catalysis
- Selectively influence a certain reaction pathway → Higher selectivity

## CO<sub>2</sub> reduction on Au in aqueous electrolyte:

CO<sub>2</sub> overtone @ 2,8 μm

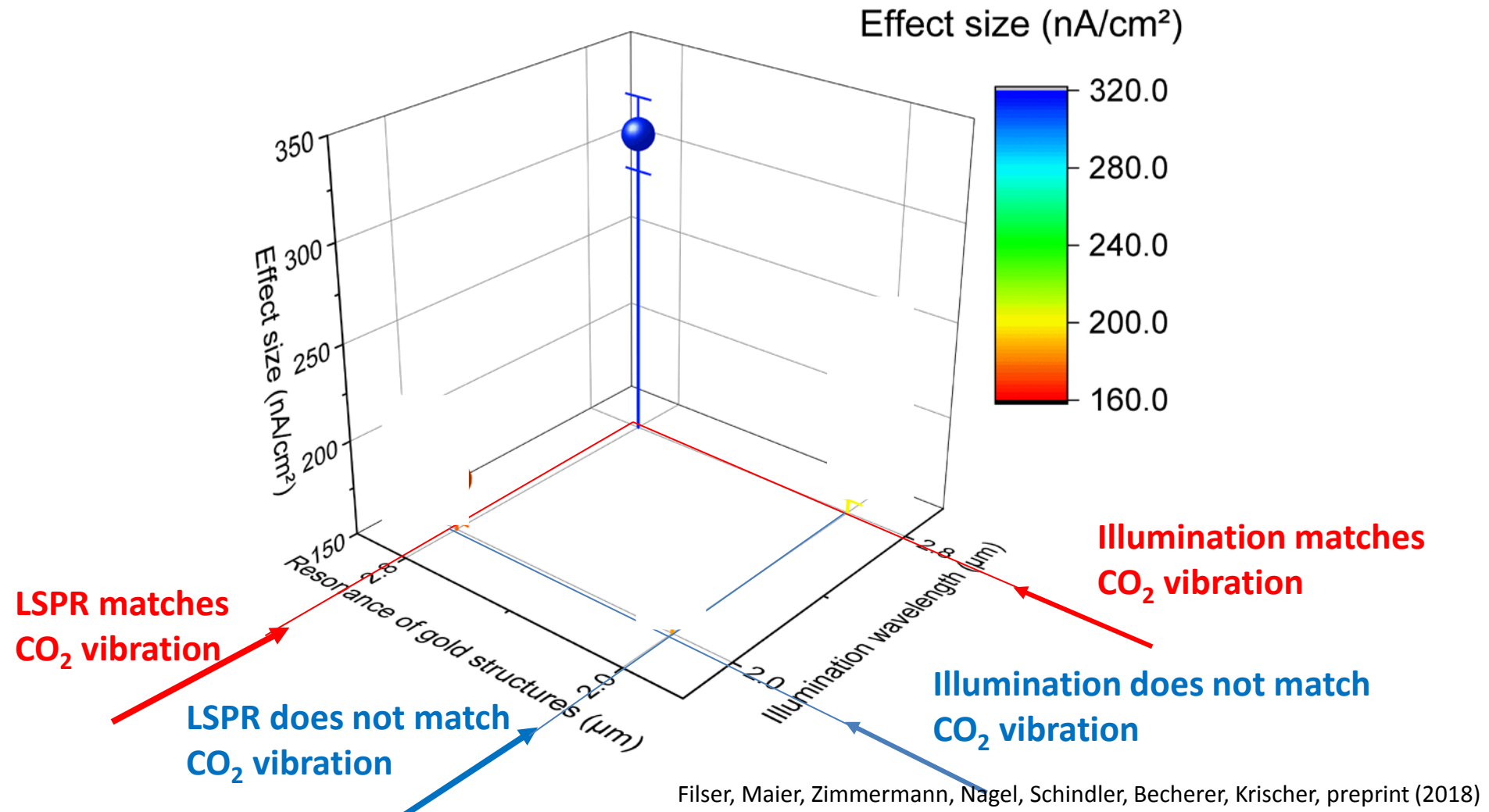


Simulation of Au nanodisks on SiO<sub>2</sub>/Si

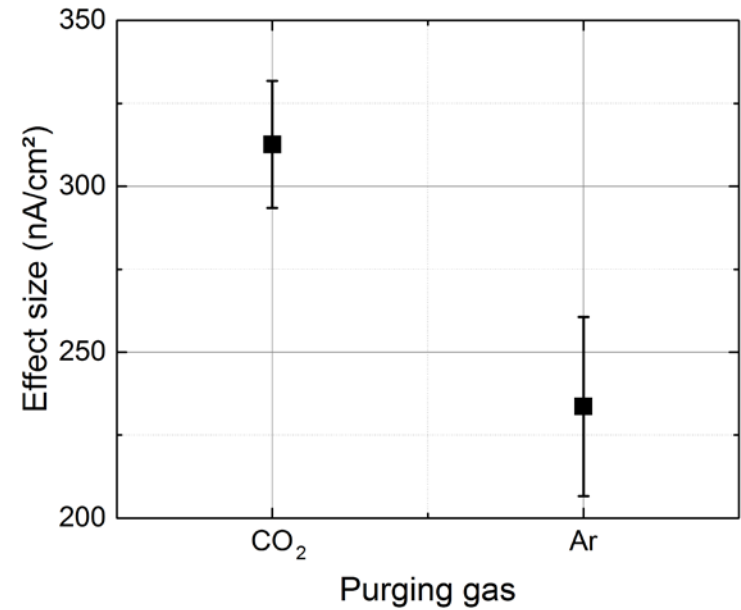
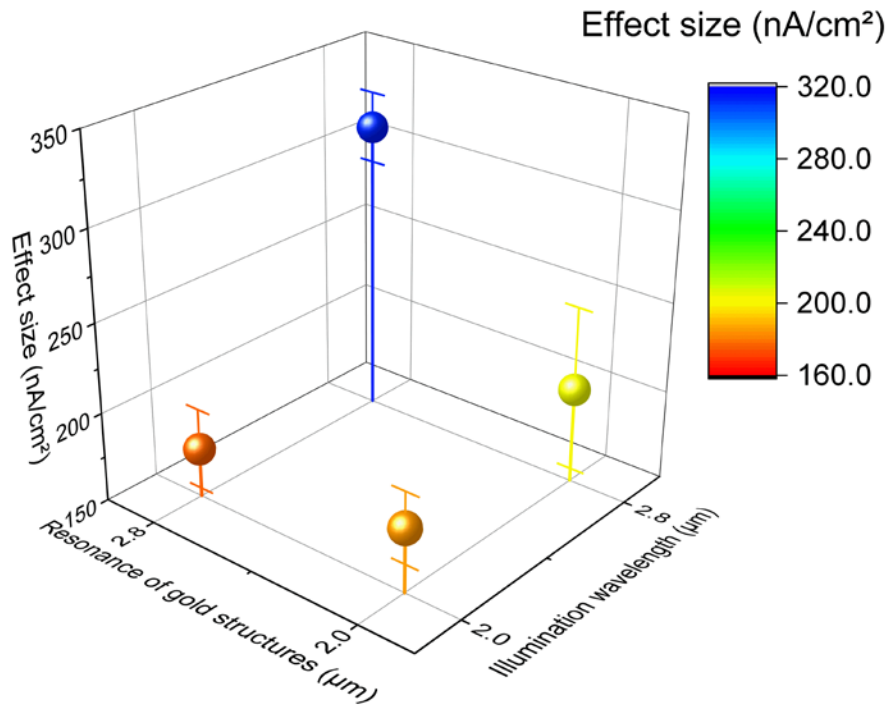


Transmission spectra of Au arrays on SiO<sub>2</sub>/Si

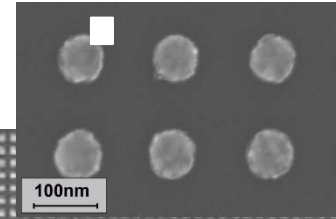
# LSPR enhanced reactivity



# ***LSPR enhanced reactivity and selectivity***



# Summary



- Au(Cu)/SiO<sub>2</sub> /p-Si MIS structures exhibit enhanced catalytic activity towards H<sub>2</sub> evolution

➔ *bifunctional mechanism / strong metal support interaction*

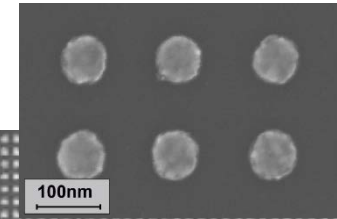
*or*

*energetics ( $\phi$  distribution at the 3 phase interface?)*

2 $\mu$ m  
|-----|

# Acknowledgements

**Simon Filser**  
**Robin Nagel (TUM EI)**  
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DFG

# *Frequency dependence of enhancement effect*

